

THE PHOTOELECTRIC EFFECT

Essential info

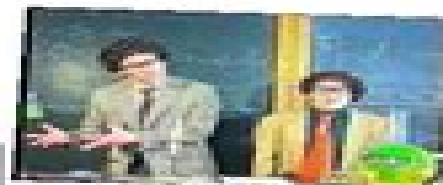
$$hf = \phi + E_{k, \max}$$

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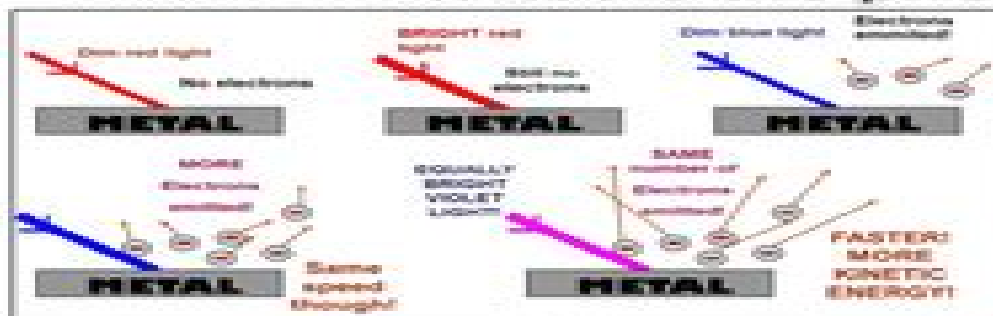
$$E = hf$$

$$1 \text{ eV} = 1.6 \times 10^{-19} \text{ J}$$

$$h = 6.63 \times 10^{-34} \text{ Js}$$



TASK 1: What is the photoelectric effect?



Different colours of light are shone on to a metal surface. The **number** of electrons emitted from the surface and their **maximum velocity** is observed.

1. Explain the observations to the left using the **photon model of light**. Use the terms "**threshold frequency**" and "**work function**" in your answers.
2. Why can't the **wave model of light** explain these observations?

TASK 2: Photoelectric Questions!

1. The **threshold frequency** of caesium metal is $4.7 \times 10^{14} \text{ Hz}$. This is in the red (visible light) section of the electromagnetic spectrum.
 - a) Explain what the above information tells us about caesium metal with respect to the **photoelectric effect**.
 - b) Calculate the **work function** of caesium metal i) in joules ii) in electron volts.
 - c) Explain if electrons will be emitted when the following types of light are shone onto the caesium metal.
 - i) infrared light is shone onto caesium.
 - ii) ultraviolet light is shone on to caesium.
 - iii) very high intensity infrared light.
 - iv) low intensity green light.
 - v) visible light of frequency $7 \times 10^{14} \text{ Hz}$.
 - vi) high intensity light of wavelength $7.4 \times 10^{-7} \text{ m}$.
2. Copper metal has a **work function** of 4.84 eV.
 - a) Calculate the **threshold frequency** of copper?
 - b) Will red light cause electrons to be emitted from this metal? (Think!)
 - c) What type of light must be shone on to this metal in order to cause electrons to be emitted?

Photoelectric Effect Gizmo Worksheet

SJ Ball



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