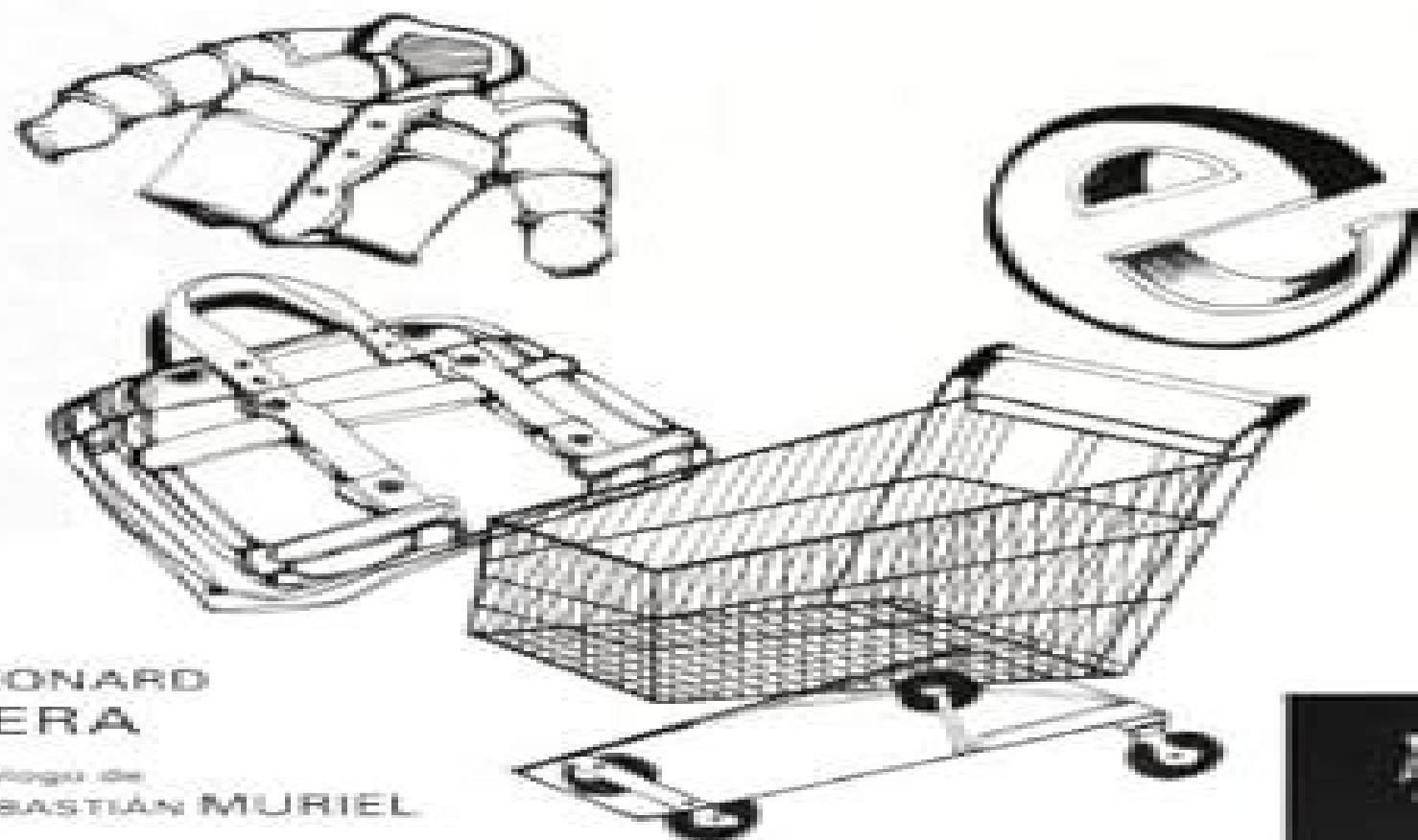


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temperature change plugging in the values given in the problem you get  $q_{\text{water}} = 4.18 \text{ J/g} \cdot \text{C} \times 110 \text{ g} \times 26.6 \text{ C} = 1220 \text{ J}$   
 $550 \text{ J} = \Delta h \cdot q_{\text{water}} = 550 \text{ J}$

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