Practical L- Borohydride reduction

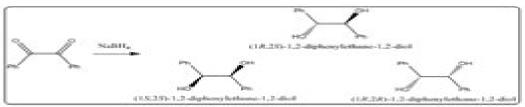


Figure 3: Reduction of Benzil to hydrobenzoin with Possible Formation of Three Stereochemical Products.

- We use sodium borohydride and methanol.
- There are two steps to the mechanism:
 - Reduction with sodium borohydride (borohydride is nucleophilic)
 - Addition of HCI -> acid workup process where you destroy any remaining borohydride. Protonation of the basic alkoxide with hydrochloric acid to from the alcohol.
- H-B-H-

- BH4 CH4 NH4+ All have 8 valence electrons therefore form a tetrahedron shape
- N(delta negative) H(delta positive). But as N is more electronegative than C, therefore -> C(delta positive) -H(delta negative)

- The phenol ring is represented on the IR spectrum which is approximately 3000cm-1.
 Therefore, when we stretch this molecule, it leads to a large spread of charge.
- If we use NaB2D4 instead of NaBH4, we get a product which has a 2D instead of H. We can then detect this using IR.
- Use symmetry in NMR to determine the environments in protons. If the molecule does not sit
 on a planar this can lead to different proton interactions, forming different environments.

Sodium Borohydride Practical Manual

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