

# Half Life and Radioactive Decay

Name: \_\_\_\_\_

**Instructions:** The *half-life* of a radioactive substance is a characteristic constant. It measures the time it takes for a given amount of the substance to become reduced by half as a consequence of decay, and therefore, the emission of radiation. Scientists use half-life to calculate amount of radioactive isotopes present in a sample.

**Sample Half-life Problem:** A 100 gram sample of carbon-14 undergoes radioactive decay until 12.5 grams of carbon-14 remains in the original sample. This process took 17,190 years. How many half-lives have passed? What is the half-life of  $^{14}\text{C}$ ?

**Step 1:** Make a data table with the original isotope, half-lives and time passed as its column headers.

Isotope (g)	Number of Half-lives Passed	Time Passed (Years)
100g	0	0
50	1	
25	2	

**Step 2:** Using the half-life principle, deduct the correct amount of isotope and record the number of half-lives passed until you reach the given number of years.

Amount of Isotope (g)	Number of Half-lives Passed	Time Passed (Years)
100g	0	

**Step 3:** Divide the number of year passed by the number of half-lives passed:  $17,190/5 = 3438$  years in one half-life. Fill in the remainder of the chart.

Amount of Isotope (g)	Number of Half-lives Passed	Time Passed (Years)
100g	0	0
50	1	3438
25	2	6876
12.5	3	10314

**Half-life Problems:**

- What is the half-life of  $^{235}\text{U}$  if a 500g sample decays to 30.25 grams in 63 hours?
- How much of a 100 g sample of  $^{14}\text{C}$  would remain after 37 years if the half-life is 12.35 years?
- The half-life of  $^{239}\text{Pu}$  is  $1.4 \times 10^4$  years. If there are 6.25 g of the sample left after  $2.8 \times 10^4$  years, how many grams were in the original sample?
- The half-life of Iodine-131 is 8.04 days. How long will it take a 200 g sample to decay into 6.25g?

# Radioactive Decay And Half Life Worksheet Answers

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