

## Nuclear Decay

### Types of Radiation

1. Complete the table below to compare the properties of alpha, beta, and gamma radiation.

Name	Alpha	Beta	Gamma
Greek letter	$\alpha$	$\beta$	$\gamma$
Symbol	${}^4_2\text{He}$	${}^0_{-1}\text{e}$	$\gamma$
Composition	2 protons, 2 neutrons	an electron	radiation (light)
Charge	0	-1	0
Stopped by	paper, skin, clothing	aluminum	lead

### Nuclear Decay Reactions

In problems 2 – 5, fill in the missing particle for each nuclear decay reaction. Then, match each equation to the appropriate type of nuclear decay.



A. alpha emission



**B. Net emissions**



C electron capture

In problems 6 and 7, write a balanced equation for each nuclear decay reaction.

6. Decay of radium-223 by alpha ( $\alpha$ ) emission



7. Decay of iodine-126 by electron ( $\beta$ ) emission



3. How are the atomic number and mass number of a nuclide affected by Beta Decay?

The mass number does not change but the atomic number increases by 1

9. How are the atomic number and mass number of a nuclide affected by the emission of gamma radiation?

The atomic number and mass number do not change

10. Which has a slightly greater mass in a nuclear reaction, the reactants or the products? Why?

It depends on the type of reaction. A fusion reaction will have greater mass in the products because it is combining two nuclei. A fission reaction will not change mass.

# Radioactive Decay Study Guide Answer Key

**Steven S. Zumdahl, Paul B. Kelter**



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