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web recall that the molar mass can be determined from a chemical formula and used as a conversion factor we can add that conversion factor as another step in a calculation to make a mole mass calculation where we start with a given number of

moles of a substance and calculate the mass of another substance involved in the chemical equation or vice

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web mass 32 g mass 16 g mass 3 2 g 2 for jars d e f use the number of molecules in the circle to determine the number of moles of O_2 in the jar then calculate the mass of the gas in the jar d e f 12 mol O_2 2 mol O_2 6 mol O_2 mass 384 g mass 64 g mass 192 g 3 jars g h and i contain ammonia gas NH_3

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web learning objective to calculate the quantities of compounds produced or consumed in a chemical reaction a balanced chemical equation gives the identity of the reactants and the products as well as the accurate number of molecules or

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web in this article we will look at how we can use the stoichiometric relationships contained in balanced chemical equations to determine amounts of substances consumed and produced in chemical reactions balanced equations and mole ratios

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web jul 4 2019 solution from the balanced equation it is known that 1 mol N_2 2 mol NH_3 use the periodic table to look at the atomic weights of the elements and calculate the weights of the reactants and products 1 mol of N_2 28.0 g 1 mol of NH_3 is 17.0 g 3 mol 51.0 g

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web solution average atomic mass of Li 6.941 amu \times 6.0151 amu 1 \times 7.0160 amu 6.941 1.0009 \times 7.0160 1.0009 \times 0.075 \times 0.075 \times 0.075 corresponds to a natural abundance of 6Li of 7.5 percent the natural abundance of 7Li is \times 0.925 or 92.5 percent 3.7 the unit factor required is 6.022

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web a molar mass Li_2CO_3 2 6.941 g 12.01 g 3 16.00 g 73.89 g b molar mass CS_2 12.01 g 2 32.07 g 76.15 g c molar mass $CHCl_3$

12 01 g 1 008 g 3 35 45 g 119 37 g d molar mass $\text{C}_6\text{H}_8\text{O}_6$ 6 12 01 g 8 1 008 g 6 16 00 g 176 12 g e molar mass KNO_3 39 10 g
 14 01 g 3 16 00 g 101 11 g

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web it is a small step from mole mass calculations to mass mass calculations if we start with a known mass of one substance in a chemical reaction instead of a known number of moles we can calculate the corresponding masses of other substances in the reaction

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web jul 4 2022 to calculate the mass of titanium metal that you can obtain multiply the number of moles of titanium by the molar mass of titanium 47 867 g mol moles ti mass ti times molar mass ti 4 12 cancel mol ti times $\frac{47\ 867\ \text{g}\ \text{ti}}{1\ \text{mol}\ \text{ti}}$ 197 g ti notag

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web 41 42 chemistry ch 3 mass relationships in chemical reactions molecular mass of NO_2 atomic mass of n 2 atomic mass of o 14 01 ainu 2 16 00 ainu 46 01 amu atomic mass of N_2O_5 2 atomic mass of n 5 atomic mass of o 2 14 01 ainu 5 16 00 ainu 108 02 amu example 3 2 exercise 3 3