

PERIODIC TABLE OF THE ELEMENTS

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PERIOD

1	1.0079 H HYDROGEN	2	4.0026 He HELIUM
3	6.941 Li LITHIUM	4	9.0122 Be BERYLLIUM
5	11.009 B BORON	6	12.011 C CARBON
7	14.007 N NITROGEN	8	15.999 O OXYGEN
9	18.998 F FLUORINE	10	20.180 Ne NEON
11	22.990 Na SODIUM	12	24.305 Mg MAGNESIUM
13	26.982 Al ALUMINUM	14	28.086 Si SILICON
15	30.974 P PHOSPHORUS	16	32.06 S SULFUR
17	35.45 Cl CHLORINE	18	39.948 Ar ARGON
19	39.098 K POTASSIUM	20	40.078 Ca CALCIUM
21	44.956 Sc SCANDIUM	22	47.867 Ti TITANIUM
23	50.942 V VANADIUM	24	51.996 Cr CHROMIUM
25	54.938 Mn MANGANESE	26	55.845 Fe IRON
27	58.933 Co COBALT	28	58.933 Ni NICKEL
29	63.546 Cu COPPER	30	65.38 Zn ZINC
31	69.723 Ga GALLIUM	32	72.64 Ge GERMANIUM
33	74.922 As ARSENIC	34	75.94 Se SELENIUM
35	78.96 Br BROMINE	36	83.798 Kr KRYPTON
37	85.468 Rb RUBIDIUM	38	87.62 Sr STRONTIUM
39	88.906 Y YTTRIUM	40	91.224 Zr ZIRCONIUM
41	92.906 Nb NIOBIUM	42	95.96 Mo MOLYBDENUM
43	98 Tc TECHNETIUM	44	101.07 Ru RUTHENIUM
45	102.91 Rh RHODIUM	46	106.42 Pd PALLADIUM
47	107.87 Ag SILVER	48	112.41 Cd CADMIUM
49	114.82 In INDIUM	50	118.71 Sn TIN
51	121.76 Sb ANTIMONY	52	127.60 Te TELLURIUM
53	126.90 I IODINE	54	131.29 Xe XENON
55	132.91 Cs CAESIUM	56	137.33 Ba BARIUM
57	138.91 La-Lu Lanthanide	58	175.48 Hf HAFNIUM
59	178.49 Ta TANTALUM	60	180.95 W TUNGSTEN
61	180.95 Re RHENIUM	62	186.21 Os OSMIUM
63	186.21 Ir IRIDIUM	64	192.22 Pt PLATINUM
65	195.08 Au GOLD	66	196.97 Hg MERCURY
67	200.59 Tl THALLIUM	68	204.38 Pb LEAD
69	207.2 Bi BISMUTH	70	208.98 Po POLONIUM
71	209 At ASTATINE	72	210 Rn RADON
73	223 Fr FRANCIUM	74	226 Ra RADIUM
75	227 Ac-Lr Actinide	76	227 Th THORIUM
77	232 Pa PROTACTINIUM	78	238 U URANIUM
79	238 Np NEPTUNIUM	80	237 Pu PLUTONIUM
81	244 Am AMERICIUM	82	243 Cm CURIUM
83	247 Bk BERKELEYIUM	84	247 Cf CALIFORNIUM
85	251 Es EINSTEINIUM	86	252 Fm FERMIUM
87	257 Md MEYERIUM	88	259 No NIOBELIUM
89	261 Lr LAWRENCIUM	90	262 Uu UNUNBIUM
91	267 Uub UNUNBIDIUM	92	269 Uut UNUNTRIUM
93	271 Uuq UNUNQUADIUM	94	272 Uup UNUNPENTIUM
95	277 Uuh UNUNHEXIUM	96	278 Uus UNUNSEPTIUM
97	283 Uuo UNUNOCTIUM	98	286 Uuq UNUNQUADIUM
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YCT Expert Team



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The Problem of Physico-chemical Periodicity Ernest Sydney Hedges, James Eckersley Myers, 1926 **A Level Chemistry MCQ (Multiple Choice Questions)** Arshad Iqbal, 2019-06-18

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Chemistry Education in the ICT Age Minu Gupta Bhowon, Sabina Jhaumeer-Laulloo, Henri Li Kam Wah, Ponnadurai Ramasami, 2009-07-21 th th The 20 International Conference on Chemical Education 20 ICCE which had rd th Chemistry in the ICT Age as the theme was held from 3 to 8 August 2008 at Le Meridien Hotel Pointe aux Piments in Mauritius With more than 200 participants from 40 countries the conference featured 140 oral and 50 poster presentations th Participants of the 20 ICCE were invited to submit full papers and the latter were subjected to peer review The selected accepted papers are collected in this book of proceedings This book of proceedings encloses 39 presentations covering topics ranging from fundamental to applied chemistry such as Arts and Chemistry Education Biochemistry and Biotechnology Chemical Education for Development Chemistry at Secondary Level Chemistry at Tertiary Level Chemistry Teacher Education Chemistry and Society Chemistry Olympiad Context Oriented Chemistry ICT and Chemistry Education Green Chemistry Micro Scale Chemistry Modern Technologies in Chemistry Education Network for Chemistry and Chemical Engineering Education Public Understanding of Chemistry Research in Chemistry Education and Science Education at Elementary Level We would like to thank those who submitted the full papers and the reviewers for their timely help in assessing the papers for publication th We would also like to pay a special tribute to all the sponsors of the 20 ICCE and in particular the Tertiary Education Commission <http://tec.intnet.mu> and the Organisation for the Prohibition of Chemical Weapons <http://www.opcw.org> for kindly agreeing to fund the publication of these proceedings EPA's Air Pollution Control Program United States. Congress. House. Committee on Energy and Commerce. Subcommittee on Oversight and Investigations, 1984 **AP Chemistry For Dummies** Peter J. Mikulecky, Michelle Rose Gilman, Kate Brutlag, 2008-11-13 A practical and hands on guide for learning the

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of the historic development of selected areas of chemistry and chemical physics discussing in detail the European American and Russian approaches to the development of chemistry Other key topics include the kinetics and non linear thermodynamics of chemical reactions and mathematical modeling which have found new applications in the theory of dynamical systems The first observations of the periodicity of chemical reactions were lost in the mist of time In the second half of the 19th century the phenomenon of chemical periodicity was studied in relation to electrochemistry solutions and colloids Discovered in the late 19th century Liesegang rings are still enigmatic and remain attractive for researchers However the discovery of the Belousov Zhabotinsky reaction marked the successful culmination of the efforts to find a true chemical oscillatory reaction The book investigates chemical phenomena that were neglected in the past but have been rediscovered placing them into a new conceptual framework For example it notes that William Bray who discovered the first oscillatory homogeneous reaction in 1921 was influenced by the first bio mathematicians who predicted chemical oscillations in homogeneous systems

Making Chemistry Relevant Sharmistha Basu-Dutt, 2010-02-19 Unique new approaches for making chemistry accessible to diverse students Students interest and achievement in academics improve dramatically when they make connections between what they are learning and the potential uses of that knowledge in the workplace and or in the world at large Making Chemistry Relevant presents a unique collection of strategies that have been used successfully in chemistry classrooms to create a learner sensitive environment that enhances academic achievement and social competence of students Rejecting rote memorization the book proposes a cognitive constructivist philosophy that casts the teacher as a facilitator helping students to construct solutions to problems Written by chemistry professors and research groups from a wide variety of colleges and universities the book offers a number of creative ways to make chemistry relevant to the student including Teaching science in the context of major life issues and STEM professions Relating chemistry to current events such as global warming pollution and terrorism Integrating science research into the undergraduate laboratory curriculum Enriching the learning experience for students with a variety of learning styles as well as accommodating the visually challenged students Using media hypermedia games and puzzles in the teaching of chemistry Both novice and experienced faculty alike will find valuable ideas ready to be applied and adapted to enhance the learning experience of all their students

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Chemistry of the Elements N. N. Greenwood,A. Earnshaw,2012-12-02 When this innovative textbook first appeared in 1984 it rapidly became a great success throughout the world and has already been translated into several European and Asian languages Now the authors have completely revised and updated the text including more than 2000 new literature references to work published since the first edition No page has been left unaltered but the novel features which proved so attractive have been retained The book presents a balanced coherent and comprehensive account of the chemistry of the elements for both undergraduate and postgraduate students This crucial central area of chemistry is full of ingenious experiments intriguing compounds and exciting new discoveries The authors specifically avoid the term inorganic chemistry since this evokes an outmoded view of chemistry which is no longer appropriate in the final decade of the 20th century Accordingly the book covers not only the inorganic chemistry of the elements but also analytical theoretical industrial organometallic bio inorganic and other cognate areas of chemistry The authors have broken with recent tradition in the teaching of their subject and adopted a new and highly successful approach based on descriptive chemistry The chemistry of the elements is still discussed within the context of an underlying theoretical framework giving cohesion and structure to the text but at all times the chemical facts are emphasized Students are invited to enter the exciting world of chemical phenomena with a sound knowledge and understanding of the subject to approach experimentation with an open mind and to assess observations reliably This is a book that students will not only value during their formal education but will keep and refer to throughout their careers as chemists Completely revised and updated Unique approach to the subject More comprehensive than competing titles American Journal of Science and Arts ,1927 **The American Journal of Science** ,1927 *The Periodic Table* I D. Michael P. Mingos,2020-02-05 As 2019 has been declared the International Year of the Periodic Table it is appropriate that Structure and Bonding marks this anniversary with two special volumes In 1869 Dmitri Ivanovitch Mendeleev first proposed his periodic table of the elements He is given the major credit for proposing the conceptual framework used by chemists to systematically inter relate the chemical properties of the elements However the concept of periodicity evolved in distinct stages and was the culmination of work by other chemists over several decades For example Newland s Law of Octaves marked an important step in the evolution of the periodic system since it represented the

first clear statement that the properties of the elements repeated after intervals of 8 Mendeleev's predictions demonstrated in an impressive manner how the periodic table could be used to predict the occurrence and properties of new elements. Not all of his many predictions proved to be valid but the discovery of scandium, gallium and germanium represented sufficient vindication of its utility and they cemented its enduring influence. Mendeleev's periodic table was based on the atomic weights of the elements and it was another 50 years before Moseley established that it was the atomic number of the elements that was the fundamental parameter and this led to the prediction of further elements. Some have suggested that the periodic table is one of the most fruitful ideas in modern science and that it is comparable to Darwin's theory of evolution by natural selection proposed at approximately the same time. There is no doubt that the periodic table occupies a central position in chemistry. In its modern form it is reproduced in most undergraduate inorganic textbooks and is present in almost every chemistry lecture room and classroom. This first volume provides chemists with an account of the historical development of the Periodic Table and an overview of how the Periodic Table has evolved over the last 150 years. It also illustrates how it has guided the research programmes of some distinguished chemists.

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In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their ability to kindle emotions, provoke contemplation, and ignite transformative change is actually awe-inspiring. Enter the realm of "**Test A Chemical Periodicity**," a mesmerizing literary masterpiece penned with a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve to the book is central themes, examine its distinctive writing style, and assess its profound effect on the souls of its readers.

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