Gases Test SCH3U Name:

___K __T __C __A Assessment of Learning (Product)
15 9 5 10

Multiple Choice. Circle the most correct answer. (15K)

1. Convert 120 K to °C.

A) 393 °C B) -153 °C C) 880 °C D) 153 °C

2. Why do we assume that gas particles experience no intermolecular forces?

A) Because they are moving too quickly

B) Because they are small.

C) Because gas particles are a long distance from each other

D) None of the above

3. The mass of a fixed quantity of a gas:

A) increases when the volume increases, at constant temperature

B) decreases when the absolute temperature decreases, at constant temperature

C) does not change at any temperature and pressure

D) decreases as the density increases

E) decreases when the volume decreases, at constant pressure

 Which rigid cylinder contains the same number of gas molecules at STP as a 2.0- liter rigid cylinder containing H₂ (g) at STP?

A) 2.0- L cylinder of CH, (g)

B) 1.5- L cylinder of NH₁ (g)

C) 4.0- L cylinder of He (g)

D) 1.0- L cylinder of O₂ (g)

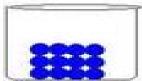
5. Which statement describes the particles of an ideal gas based on the kinetic molecular theory?

A) The gas particles are relatively far apart and have negligible volume.

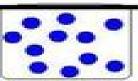
B) The gas particles are in constant, nonlinear motion.

C) The gas particles have attractive forces between them.

D) When gas particles have collisions they lose energy.







6. Which phase depicted above does not have a definite shape or volume?

A) The one in the middle

B) The one on the left

C) The one in the middle and the one on the right

D) The one on the right

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Ying-Ying Zheng

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- 1. Understanding the eBook Test Questions And Answers For The Nature Of Gases
 - The Rise of Digital Reading Test Questions And Answers For The Nature Of Gases
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Test Questions And Answers For The Nature Of Gases
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Test Questions And Answers For The Nature Of Gases
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Test Questions And Answers For The Nature Of Gases
 - Personalized Recommendations
 - Test Questions And Answers For The Nature Of Gases User Reviews and Ratings
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Test Questions And Answers For The Nature Of Gases
- 9. Balancing eBooks and Physical Books Test Questions And Answers For The Nature Of Gases
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 - o Creating a Diverse Reading Collection Test Questions And Answers For The Nature Of Gases
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 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Test Questions And Answers For The Nature Of Gases
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Test Questions And Answers For The Nature Of Gases
 - Fact-Checking eBook Content of Test Questions And Answers For The Nature Of Gases
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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