

## Molecular (Covalent) Formula Writing

Carbon = 4	Chlorine = 7	Hydrogen = 1	Oxygen = 6	Sulfur = 6
Barium = 20	Argon = 18	Calcium = 2	Fluorine = 7	Neon = 10

Remember: though, ionic compounds cannot share electrons, a stable compound can form through the sharing of electrons among the molecules it forms. After molecular formulas are used for determining the chemical formula of a compound, bond pairs are written. Compounds that contain ionic compounds are written. When an ionic compound is written, the first symbol should be capital. Remember, always use the correct formula and a line between the two symbols to indicate the bond pair. Remember, always use the correct formula and a line between the two symbols to indicate the bond pair. Remember, always use the correct formula and a line between the two symbols to indicate the bond pair.



Write the chemical formula for the following compounds:

1. Carbon monoxide  
C O  
CO

Write the chemical formula for the following compounds:

Compound Name	Chemical Formula	Compound Name	Chemical Formula
Hydrogen chloride	<u>H</u> <u>Cl</u> <u>HCl</u>	Hydrogen chloride	<u>H</u> <u>Cl</u> <u>HCl</u>
Carbon monoxide	<u>C</u> <u>O</u> <u>CO</u>	Carbon monoxide	<u>C</u> <u>O</u> <u>CO</u>
Carbon dioxide	<u>C</u> <u>O</u> <u>O</u> <u>CO2</u>	Carbon dioxide	<u>C</u> <u>O</u> <u>O</u> <u>CO2</u>
Hydrogen peroxide	<u>H</u> <u>O</u> <u>O</u> <u>H2O2</u>	Hydrogen peroxide	<u>H</u> <u>O</u> <u>O</u> <u>H2O2</u>
Carbon tetrachloride	<u>C</u> <u>Cl</u> <u>Cl</u> <u>Cl</u> <u>Cl</u> <u>CCl4</u>	Carbon tetrachloride	<u>C</u> <u>Cl</u> <u>Cl</u> <u>Cl</u> <u>Cl</u> <u>CCl4</u>
Sulfur dioxide	<u>S</u> <u>O</u> <u>O</u> <u>SO2</u>	Sulfur dioxide	<u>S</u> <u>O</u> <u>O</u> <u>SO2</u>
Phosphorus pentachloride	<u>P</u> <u>Cl</u> <u>Cl</u> <u>Cl</u> <u>Cl</u> <u>Cl</u> <u>PCl5</u>	Phosphorus pentachloride	<u>P</u> <u>Cl</u> <u>Cl</u> <u>Cl</u> <u>Cl</u> <u>Cl</u> <u>PCl5</u>

# Test Unit 10 Bonding And Formula Writing Answers

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