

NUCLEAR CHEMISTRY AND APPLICATIONS OF RADIOACTIVITY

Nuclear chemistry: The branch of chemistry which deals with the study of radioactivity, nuclear processes and transformations in the nuclei of atoms etc. is called nuclear chemistry.

Radioactivity: The phenomenon of spontaneous emission of invisible radiations like α -particles, β -particles and γ -particles from unstable nuclei to get the stability is called radioactivity.

The substances which show radioactivity are called radioactive substances. For example: Uranium, thorium, polonium etc.

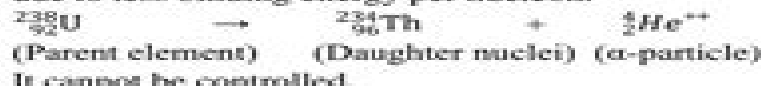
Units of radioactivity:

The unit of radioactivity is Curie (Ci) and defined as the mass of radioactive substance that can produce 3.7×10^{10} dis-integration per second. Its other units are milli curie and micro curie.

$$1 \text{ Ci} = 3.7 \times 10^{10} \text{ dps} = 3.7 \times 10^{10} \text{ Bq}$$

Note: Bq (Becquerel) is another unit of radioactivity.

Natural radioactivity: The process of self disintegration of unstable nuclei by emission of different types of radiations to get stability is called natural radioactivity. For example: Uranium, polonium and radium show natural radioactivity. Generally, heavier nuclei show natural radioactivity due to less binding energy per nucleon.

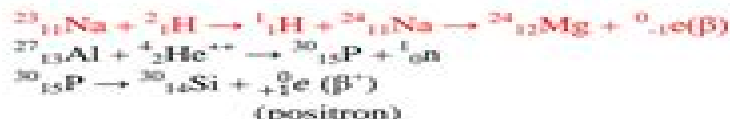


It cannot be controlled.

Artificial radioactivity (Transmutation): The process by which an element gets converted into new element by bombarding it with high energy particles is called artificial radioactivity.

It can be controlled by controlling the speed of bombarding particles.

For example:



Nuclear reactions: The reaction in which the compositions of certain nuclei undergo a change are called nuclear reactions.

There are two types of nuclear reactions and they are:

1. Nuclear fission reaction
2. Nuclear fusion reaction

Nuclear fission reaction: The nuclear reaction in which a heavy nucleus splits into two or more lighter nuclei with the simultaneous emission of large amount of energy is called nuclear fission reaction.

A chain reaction is set up in nuclear fission reaction. The products of the reaction are radioactive. For example:



Nuclear fusion reaction: The nuclear reaction in which two lighter nuclei combine with each other to give stable nuclei with higher mass is called nuclear fusion.

They are generally carried out at very high temperature. A chain reaction is not set up in the nuclear reaction. The products of the reaction are not radioactive. For example:



The total energy produced during formation of one nuclei of helium is 26.82MeV.

Note: A practical application of nuclear fusion reaction is the production of hydrogen bomb, which is the misuse of nuclear power. Hydrogen bomb is 1000 times powerful than the atomic bomb.

Differences between nuclear fission and nuclear fusion reactions: The differences between nuclear fission and nuclear fusion reactions:

Nuclear fission reaction	Nuclear fusion reaction
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Radioactivity And Nuclear Chemistry Answers

**United States. National Bureau of
Standards, Betty L. Oberholtzer**



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Nuclei Isotopes and Isotope Separation
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Radionuclides in Nature
Absorption of Nuclear Radiation
Radiation Effects on Matter
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Radiochemistry
Karl Heinrich Lieser, 2008-09-26
This handbook gives a complete and concise description of the up to date knowledge of nuclear and radiochemistry and applications in the various fields of science. It is based on teaching courses and on research for over 40 years. The book is addressed to any researcher wishing sound knowledge about the properties of matter be it a chemist, a physicist, a medical doctor, a mineralogist or a biologist. They will all find it a valuable source of information about the principles and applications of nuclear and radiochemistry. Research in radiochemistry includes: Study of radioactive matter in nature, investigation of radioactive transmutations by chemical methods, chemistry of radioelements etc. Applications include: Radionuclides in geo and cosmochemistry, dating by nuclear methods, radioanalysis, Mossbauer spectroscopy and related methods, behaviour of natural and man made radionuclides in the environment, dosimetry and radiation protection. All subjects are presented clearly and comprehensibly and in logical sequence. Detailed derivations of equations are avoided and relevant information is compiled in tables. The recent edition of the multi coloured Karlsruhe Chart of the Nuclides is included. Clearly a standard work by an author with extensive experience in research and teaching. *The Handy Chemistry Answer Book*
Justin P. Lomont, Ian C. Stewart, 2013-10-01
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