

ANSWERS TO UNIT V : ELECTROCHEMISTRY

1. (a) $\text{Mn} \longrightarrow \text{Mn}^{2+} + 2\text{e}^-$; Mn is oxidized and is the reducing agent.
 $\text{Hg}^{2+} + 2\text{e}^- \longrightarrow \text{Hg}$; Hg^{2+} is reduced and is the oxidizing agent.
- (b) $\text{H}_2 \longrightarrow 2\text{H}^+ + 2\text{e}^-$; H_2 is oxidized and is the reducing agent.
 $\text{Sn}^{4+} + 2\text{e}^- \longrightarrow \text{Sn}^{2+}$; Sn^{4+} is reduced and is the oxidizing agent.
- (c) $\text{Li} \longrightarrow \text{Li}^+ + \text{e}^-$; Li is oxidized and is the reducing agent.
 $\text{F}_2 + 2\text{e}^- \longrightarrow 2\text{F}^-$; F_2 is reduced and is the oxidizing agent.
- (d) $\text{Cr}^{2+} \longrightarrow \text{Cr}^{3+} + \text{e}^-$; Cr^{2+} is oxidized and is the reducing agent.
 $\text{Br}_2 + 2\text{e}^- \longrightarrow 2\text{Br}^-$; Br_2 is reduced and is the oxidizing agent.
- (e) $\text{Fe}^{2+} \longrightarrow \text{Fe}^{3+} + \text{e}^-$; Fe^{2+} is oxidized and is the reducing agent.
 $\text{Sn}^{4+} + 2\text{e}^- \longrightarrow \text{Sn}^{2+}$; Sn^{4+} is reduced and is the oxidizing agent.
2. (a) $2\text{Cs} + \text{Cl}_2 \longrightarrow 2\text{CsCl}$
- (b) $\text{Cs} \longrightarrow \text{Cs}^+ + \text{e}^-$ (oxidation) $\text{Cl}_2 + 2\text{e}^- \longrightarrow 2\text{Cl}^-$ (reduction)
- (c) Cs is the reducing agent; Cl_2 is the oxidizing agent.
3. (a) +5 (d) +6 (g) -3 (j) +5 (m) +5 (o) -1 (q) +6 (s) 0
 (b) +3 (e) -3 (h) -8/3 (k) +3 (n) -2 (p) +3 (r) -4/3 (t) -3/2
 (c) +6 (f) -1/3 (i) +3 (l) +7
4. a) $\text{ClO}_2 + \text{C} \longrightarrow \text{ClO}_2^- + \text{CO}_3^{2-}$; C is oxidized
 +4 0 +3 +4
- b) $\text{Sn}^{2+} + \text{Cl}^- + \text{BrO}_3^- \longrightarrow \text{SnCl}_6^{2-} + \text{Br}^-$; Sn is oxidized
 +2 +5 +4 -1
- c) $\text{MnO}_4^- + \text{C}_2\text{O}_4^{2-} \longrightarrow \text{MnO}_2 + \text{CO}_2$; C is oxidized
 +7 +3 +4 +4
- d) $\text{NO}_3^- + \text{H}_2\text{Te} \longrightarrow \text{NO} + \text{TeO}_4^{2-}$; Te is oxidized
 +5 -2 +2 +6
5. Equations a, d and f
6. (a) Cl_2 , Cl^- , Cl_2O (b) NO_3^- , NO_2^- , N_2O_3
7. (a) Na^+ can only be reduced (g) Fe^{2+} can be either reduced or oxidized
 (b) F^- can only be oxidized (h) Co^{2+} can only be reduced
 (c) Cu^+ can be either reduced or oxidized (i) Se(s) cannot react at all (needs H^+ to react)
 (d) Sn^{4+} can only be reduced (j) Sn^{2+} can be either reduced or oxidized
 (e) NO_3^- cannot react at all (needs H^+ to react) (k) Al(s) can only be oxidized
 (f) Hg(l) can only be oxidized (l) acidic $\text{Cr}_2\text{O}_7^{2-}$ can only be reduced
8. (a) no reaction; the species to be reduced (Ni^{2+}) is below the species to be oxidized (Ag) on the Table.
 (b) spontaneous: $2\text{Li} + \text{Zn}^{2+} \longrightarrow 2\text{Li}^+ + \text{Zn}$.
 (c) no reaction; both species can only be oxidized.
 (d) no reaction; the species to be reduced (H^+) is below the species to be oxidized (Cu) on the Table.
 (e) spontaneous: $2\text{H}^+ + \text{Fe} \longrightarrow \text{H}_2 + \text{Fe}^{2+}$.

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