

Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The charge on an ion is
a. always positive.
b. always negative.
c. either positive or negative.
d. zero.
- _____ 2. According to the octet rule, a calcium atom has a tendency to
a. lose one electron.
b. lose two electrons.
c. gain one electron.
d. gain two electrons.
- _____ 3. If a compound forms by ionic bonding, which is *not* true?
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.
b. The net charge of the compound is zero.
c. The compound contains just two atoms, each of opposite charge.
d. Several ions group together in a tightly packed structure.
- _____ 4. The only property listed that is *not* characteristic of ionic compounds is
a. high melting point.
b. hardness.
c. lack of crystal structure.
d. good conductor of electricity.
- _____ 5. Which formula listed below represents a polyatomic ion?
a. NH_4^+
b. H_2SO_4
c. NaCl
d. H_2O
- _____ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because
a. ionic substances tend to vaporize at room temperature.
b. ionic substances are brittle.
c. attractive forces between ions are greater than the attractive forces between molecules.
d. the numbers of positive and negative charges are equal in an ionic compound.

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M Lipman



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Table of Contents Test B Ch 6 Chemical Bonds Holt

1. Understanding the eBook Test B Ch 6 Chemical Bonds Holt
 - The Rise of Digital Reading Test B Ch 6 Chemical Bonds Holt
 - Advantages of eBooks Over Traditional Books
2. Identifying Test B Ch 6 Chemical Bonds Holt
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Test B Ch 6 Chemical Bonds Holt
 - User-Friendly Interface
4. Exploring eBook Recommendations from Test B Ch 6 Chemical Bonds Holt
 - Personalized Recommendations
 - Test B Ch 6 Chemical Bonds Holt User Reviews and Ratings
 - Test B Ch 6 Chemical Bonds Holt and Bestseller Lists
5. Accessing Test B Ch 6 Chemical Bonds Holt Free and Paid eBooks
 - Test B Ch 6 Chemical Bonds Holt Public Domain eBooks
 - Test B Ch 6 Chemical Bonds Holt eBook Subscription Services
 - Test B Ch 6 Chemical Bonds Holt Budget-Friendly Options

6. Navigating Test B Ch 6 Chemical Bonds Holt eBook Formats
 - ePub, PDF, MOBI, and More
 - Test B Ch 6 Chemical Bonds Holt Compatibility with Devices
 - Test B Ch 6 Chemical Bonds Holt Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Test B Ch 6 Chemical Bonds Holt
 - Highlighting and Note-Taking Test B Ch 6 Chemical Bonds Holt
 - Interactive Elements Test B Ch 6 Chemical Bonds Holt
8. Staying Engaged with Test B Ch 6 Chemical Bonds Holt
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Test B Ch 6 Chemical Bonds Holt
9. Balancing eBooks and Physical Books Test B Ch 6 Chemical Bonds Holt
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Test B Ch 6 Chemical Bonds Holt
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Test B Ch 6 Chemical Bonds Holt
 - Setting Reading Goals Test B Ch 6 Chemical Bonds Holt
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Test B Ch 6 Chemical Bonds Holt
 - Fact-Checking eBook Content of Test B Ch 6 Chemical Bonds Holt
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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