

1. Pressure has an appreciable effect on the solubility of \_\_\_\_\_ in liquids.

- A. gases
- B. solids
- C. liquids
- D. salts
- E. solids and liquids

A

2. A saturated solution \_\_\_\_\_

- A. contains as much solvent as it can hold
- B. contains no double bonds
- C. contains dissolved solute in equilibrium with undissolved solute
- D. will rapidly precipitate if a seed crystal is added
- E. cannot be attained

C

3. A 0.100 m solution of which one of the following solutes will have the highest vapor pressure?

- A.  $\text{KClO}_4$
- B.  $\text{Ca}(\text{ClO}_4)_2$
- C.  $\text{Al}(\text{ClO}_4)_3$
- D. sucrose
- E.  $\text{NaCl}$

D

4. Which of the following liquids will have the highest freezing point?

- A. pure  $\text{H}_2\text{O}$
- B. aqueous glucose (0.60 m)
- C. aqueous sucrose (0.60 m)
- D. aqueous  $\text{FeI}_3$  (0.24 m)
- E. aqueous  $\text{KF}$  (0.50 m)

A

5. Calculate the vapor pressure of a solution made by dissolving 109 grams of glucose (molar mass = 180.2 g/mol) in 920.0 mL of water at 25°C. The vapor pressure of pure water at 25°C is 23.76 mm Hg. Assume the density of the solution is 1.00 g/mL.

- A. 0.278 mm Hg
- B. 0.605 mm Hg
- C. 22.98 mm Hg
- D. 23.48 mm Hg
- E. 23.76 mm Hg

D

# Texas Eoc Chemistry Practice Test 2

**David Brown, Sue Jones**



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