Name:

#### Neutralization Reactions Worksheet Reading: Cycle 6 Chemistry II Lesson 4

Neutralization Reaction: Acid + Base → Salt + Water
H\* from acid and OH\* from base combine to form H<sub>2</sub>O. Leftovers form 'salt'.
Remember in salt, write metal first! KCl not CIK!

Example 1:  $H + N_0OH \rightarrow H_2O + N_0I$ 

H and OH combine to form water. One atom of I & one of Na leftover. Na is a metal so write NaI (not INa)

Example 2: \_\_HF + \_ Ca(OH)<sub>2</sub> -- \_\_\_ + \_\_\_

For complete neutralization, must have equal number of H and OH. Therefore we need 2 HF for 1 Ca(OH)<sub>2</sub>

\_2\_HF + \_1\_Ca(OH)<sub>2</sub> -- \_\_\_\_ +\_\_\_\_

Two H and two OH make two waters. Think of prom! Two boys and two girls make two couples.

\_2\_HF + \_1\_Ca(OH)<sub>2</sub> → 2 H<sub>2</sub>O +

What's left? 2 F atoms and 1 Ca atom form the 'salt'. Remember that Ca is the metal, so it goes first.

 $2 \text{ HF} + 1 \text{ Ca(OH)}_2 - 2 \text{ H}_2 \text{O} + \text{CaF}_2$ 

Safety check: is CaF<sub>2</sub> a good formula? P.T. tells you ions are Ca<sup>rd</sup> and F, so total charge is 0 like it should be.

Complete & balance these neutralization reactions. Remember that H<sub>2</sub>O = HOH!

 $H_2SO_4 + Sr(OH)_2 \rightarrow __H_2O + SrSO_4$   $_Ca(OH)_2 + __H_3PO_4 \rightarrow __H_2O + __Ca_5(PO_4)_2$   $_HBr + Ba(OH)_2 \rightarrow __H_2O +$ 

 $Zn(OH)_2 + _HNO_3 - __ + Zn(NO_3)_2$ 

Al(OH)<sub>3</sub> + \_\_HCl -- \_\_\_\_ + \_\_\_\_

Complete and balance the following equations representing neutralization reactions:

a) \_CsOH + H<sub>2</sub>SO<sub>3</sub> - \_H<sub>2</sub>O + Cs<sub>2</sub>CO<sub>3</sub>

b) \_HF + Mg(OH)<sub>2</sub> -- \_H<sub>2</sub>O + \_\_\_\_\_

c)  $\_HNO_3$  +  $Al(OH)_3$  -  $\_H_2O$  +  $Al(NO_3)_3$ 

d) HCl + KOH - \_\_\_\_ + \_\_\_\_\_

e) HBrO<sub>3</sub> + LiOH - H<sub>2</sub>O + \_\_\_\_\_

## **Neutralization Reactions Work Sheet 15**

**D** Siedentop

**Neutralization Reactions Work Sheet 15:** 

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