

Name: _____

Neutralization Reactions Worksheet

Reading: Cycle 6 Chemistry II Lesson 4

Neutralization Reaction: Acid + Base \rightarrow Salt + Water

H⁺ from acid and OH⁻ from base combine to form H₂O. Leftovers form 'salt'.

Remember in salt, write metal first! KCl not ClK!

Example 1: $\text{HI} + \text{NaOH} \rightarrow \text{H}_2\text{O} + \text{NaI}$

H and OH combine to form water. One atom of I & one of Na leftover. Na is a metal so write NaI (not INa)

Example 2: $\underline{\hspace{1cm}} \text{HF} + \underline{\hspace{1cm}} \text{Ca(OH)}_2 \rightarrow \underline{\hspace{2cm}} + \underline{\hspace{2cm}}$

For complete neutralization, must have equal number of H and OH. Therefore we need 2 HF for 1 Ca(OH)₂

$\underline{\hspace{1cm}} \text{2} \text{HF} + \underline{\hspace{1cm}} \text{1} \text{Ca(OH)}_2 \rightarrow \underline{\hspace{2cm}} + \underline{\hspace{2cm}}$

Two H and two OH make two waters. Think of prom! Two boys and two girls make two couples.

$\underline{\hspace{1cm}} \text{2} \text{HF} + \underline{\hspace{1cm}} \text{1} \text{Ca(OH)}_2 \rightarrow \underline{\hspace{1cm}} \text{2 H}_2\text{O} + \underline{\hspace{2cm}}$

What's left? 2 F atoms and 1 Ca atom form the 'salt'. Remember that Ca is the metal, so it goes first.

$\underline{\hspace{1cm}} \text{2} \text{HF} + \underline{\hspace{1cm}} \text{1} \text{Ca(OH)}_2 \rightarrow \underline{\hspace{1cm}} \text{2 H}_2\text{O} + \underline{\hspace{1cm}} \text{CaF}_2$

Safety check: is CaF₂ a good formula? P.T. tells you ions are Ca⁺² and F⁻, so total charge is 0 like it should be.

1. Complete & balance these neutralization reactions. Remember that H₂O = HOH!



2. Complete and balance the following equations representing neutralization reactions:



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Clyde Metz



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