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 $\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow \text{Na}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3 + \text{NaBr}$
 $\text{SNO}_2 + \text{NF}_3 + \text{SnF}_2 + \text{N}_2\text{O}_3$
 using the following equation
 $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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 $\text{SiO}_2 + 3\text{C} \rightarrow \text{SiC} + 2\text{CO}$

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web balancing equations and simple stoichiometry answers balance the following equations $1\text{ N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$ $2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$ $3\text{HBr} + 1\text{KHCO}_3 \rightarrow 1\text{H}_2\text{O} + 1\text{KBr} + 1\text{CO}_2$ $4\text{Ga} + 3\text{Na}_2\text{SO}_3 \rightarrow 2\text{Ga}_2\text{SO}_3 + 3\text{Na}_2\text{S}$ $5\text{Sn} + 3\text{NO}_3^- \rightarrow 3\text{SnF}_2 + 1\text{N}_2\text{O}_3$

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web google classroom balance the following chemical equation $\text{Mg} + \text{OH}^- \rightarrow \text{Mg}(\text{OH})_2$ $\text{HCl} + \text{MgCl}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}$ note all reactants and products require a coefficient of at least one stuck review related articles videos or use a hint report a problem $7.4 \times 10^5 \theta \pi 8.5 \times 10^6$

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web a common type of stoichiometric relationship is the mole ratio which relates the amounts in moles of any two substances in a chemical reaction we can write a mole ratio for a pair of substances by looking at the coefficients in front of each species in

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