

Thin Layer Chromatography

I. Purpose

Thin Layer Chromatography (TLC) is a laboratory technique used to identify different compounds present within a mixture. There are two phases: The stationary phase, which contains a solid and usually polar, and the mobile phase, which can be either a liquid or gas which can be a pure solvent or a mixture of solvents.¹ The TLC plates consist of a solid surface such as glass, metal or plastic with a thin adsorbent later on the surface and provides the stationary phase. In the lab, silica gel will be act as the solid, stationary phase and 99:1 ethyl acetate. The liquid, mobile phase will be represented by the use of acetic acid. If the compound in question forms dark spots on the TLC plate, then it will stick to the stationary phase because it is polar. They tend to form spots close to the starting, origin line of the TLC plate. On the other hand, the spots will flow with the mobile phase because the compound is less polar compared to the other compound. Nonpolar substances will also form spots that traveled a further distance compared to those that are slightly polar. A more polar compound will tend to strongly bind to the silica gel compared a compound less polar. To compare the distances between each substance, a retention factor (R_f) is calculated for each spot. It is determined by the ratio traveled by the compound divided to the distance the solvent has traveled.¹ An R_f value can be enlarged by raising the polarity of the solvent, while it can also be decreased by increasing the polarity of the compound.²

II. Safety

Safety Data Sheet – NFPA ^{3,4}				
	Acetaminophen ⁵	Aspirin/Acetylsalicylic Acid ⁶	Dichloromethane ⁷	Ethyl Alcohol ⁸
Health	2	1	2	3
Flammability	1	1	1	2
Corrosivity	0	1	0	0
Other Hazards	N/A	0	N/A	0

Table 1: Safety Data Sheet

III. Experiment⁴

1. Obtain a TLC Plate of the appropriate size that can fit into the developing chamber.

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Terry C. Jones



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Table of Contents Paper Chromatography Lab Report Discussion

1. Understanding the eBook Paper Chromatography Lab Report Discussion
 - The Rise of Digital Reading Paper Chromatography Lab Report Discussion
 - Advantages of eBooks Over Traditional Books
2. Identifying Paper Chromatography Lab Report Discussion
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Paper Chromatography Lab Report Discussion
 - User-Friendly Interface
4. Exploring eBook Recommendations from Paper Chromatography Lab Report Discussion
 - Personalized Recommendations
 - Paper Chromatography Lab Report Discussion User Reviews and Ratings
 - Paper Chromatography Lab Report Discussion and Bestseller Lists
5. Accessing Paper Chromatography Lab Report Discussion Free and Paid eBooks

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 - Paper Chromatography Lab Report Discussion eBook Subscription Services
 - Paper Chromatography Lab Report Discussion Budget-Friendly Options
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 - Paper Chromatography Lab Report Discussion Compatibility with Devices
 - Paper Chromatography Lab Report Discussion Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Paper Chromatography Lab Report Discussion
 - Highlighting and Note-Taking Paper Chromatography Lab Report Discussion
 - Interactive Elements Paper Chromatography Lab Report Discussion
 8. Staying Engaged with Paper Chromatography Lab Report Discussion
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Paper Chromatography Lab Report Discussion
 9. Balancing eBooks and Physical Books Paper Chromatography Lab Report Discussion
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Paper Chromatography Lab Report Discussion
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Paper Chromatography Lab Report Discussion
 - Setting Reading Goals Paper Chromatography Lab Report Discussion
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Paper Chromatography Lab Report Discussion
 - Fact-Checking eBook Content of Paper Chromatography Lab Report Discussion
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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