

2 (a) State briefly what is meant by the following terms.

(i) reversible reaction

.....

(ii) dynamic equilibrium

.....

.....

[2]

(b) Water ionises to a small extent as follows.



(i) Write an expression for K_c for this reaction.

.....

(ii) Write down the expression for K_w , the ionic product of water, and explain how this can be derived from your K_c expression in (i).

.....

.....

(iii) State and explain how the value of K_w for hot water will differ from its value for cold water.

.....

.....

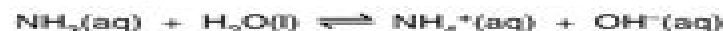
[3]

(c) K_w can be used to calculate the pH of solutions of strong and weak bases.

(i) Use the value of K_w in the *Data Booklet* to calculate the pH of $0.050 \text{ mol dm}^{-3}$ NaOH.

pH =

Ammonia ionises slightly in water as follows.



The following expression applies to this equilibrium.

$$[\text{H}_2\text{O}] \times K_c = [\text{NH}_4^+][\text{OH}^-]/[\text{NH}_3] = 1.8 \times 10^{-5} \text{ mol dm}^{-3}$$

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