

Form Code A

NAME _____

CHM 2045, Fall 2018

Final Exam Review (Sumner, Gower, Korolev, Angerhofer, Polanco)

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1. A 1.0 g organic compound sample produced 2.36 grams of CO_2 and 0.640 grams of H_2O during combustion analysis. If the hydrocarbon has a molar mass between 110 and 114 g/mol, what is its molecular formula?

- (1) C_3H_8 (2) $\text{C}_3\text{H}_8\text{O}$ (3) $\text{C}_4\text{H}_8\text{O}_2$ (4) CH_2O_2 (5) C_4H_8

2. A 6.00 M solution of dye is diluted by taking 13.00 ml of it and adding enough water to make 100.0 ml solution. Then, 18.00 ml of that solution is diluted to a volume of 500.0 ml. What is the concentration of the diluted dye?

- (1) 0.780 M (2) 0.0281 M (3) 0.0400 M (4) 0.00281 M
(5) 0.0780 M

3. Which of the following is the net ionic equation for the reaction between $\text{Ba}(\text{OH})_2$ and CuSO_4 ?

- (1) $\text{Ba}(\text{OH})_2 (\text{aq}) + \text{CuSO}_4 (\text{aq}) \rightarrow \text{BaSO}_4 (\text{s}) + \text{Cu}(\text{OH})_2 (\text{s})$
(2) $\text{Ba}(\text{OH})_2 (\text{aq}) + \text{CuSO}_4 (\text{aq}) \rightarrow \text{Ba}^{2+} (\text{aq}) + 2\text{OH}^- (\text{aq}) + \text{Cu}^{2+} (\text{aq}) + \text{SO}_4^{2-} (\text{aq})$
(3) $\text{Ba}^{2+} (\text{aq}) + 2\text{OH}^- (\text{aq}) + \text{Cu}^{2+} (\text{aq}) + \text{SO}_4^{2-} (\text{aq}) \rightarrow \text{BaSO}_4 (\text{s}) + \text{Cu}(\text{OH})_2 (\text{s})$
(4) $\text{Ba}^{2+} (\text{aq}) + \text{SO}_4^{2-} (\text{aq}) \rightarrow \text{BaSO}_4 (\text{s})$
(5) $2\text{OH}^- (\text{aq}) + \text{Cu}^{2+} (\text{aq}) \rightarrow \text{Cu}(\text{OH})_2 (\text{s})$

4. How many grams of precipitate will form if excess KCl is added to 1.46 L of 0.913 M AgNO_3 ?

- (1) 1.33g (2) 1.91×10^2 g (3) 2.30×10^2 g (4) 3.00×10^2 g
(5) 6.90×10^2 g

5. The density of fresh milk (as it comes from a yak) is found to be 83.65 lbs per cubic foot. Converted to SI units this translates into:

- (1) 83.54 kg/m^3 (2) 1997.0 kg/m^3 (3) 1028 kg/m^3 (4) 1340 kg/m^3
(5) 4998 kg/m^3

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Bici Hendricks



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