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SECTION A

1. a) A compound Q contains 64.9% carbon, 13.5% hydrogen and the rest being oxygen. 1.85g of Q in the vapour form occupied 969.8cm^3 at 200°C .

- Calculate the empirical formula of Q
- Determine the molecular formula of Q

(The molar gas constant, $R = 8.31\text{JK}^{-1}\text{mol}^{-1}$)

- Q reacts with sodium with effervescence but has no effect on sodium carbonate. Write the names and the structures of all possible isomers of Q
- When treated with anhydrous zinc chloride in the presence of concentrated hydrochloric acid, Q formed two layers after about 8 minutes. Identify Q.
- Q reacted with acidified chromium trioxide to give a compound A. Write equation for the reaction:
 - leading to the formation of A
 - between A and acidified 2, 4-dinitrophenylhydrazine and outline a mechanism for the reaction.
- Write equation to show how Q can be prepared from an alkene and outline a mechanism for the reaction.

2. a) Explain the principle on which steam distillation is based

b) State

- the conditions necessary for steam distillation
- one advantage of steam distillation over fractional distillation
- The following data was obtained for the steam distillation of bromobenzene at 760mmHg pressure

Temperature ($^\circ\text{C}$)	90	92	94	96	98	100
Vapour pressure of water (mmHg)	526	567	611	658	707	760
Vapour pressure of bromobenzene (mmHg)	96	106	114	123	132	141

- On the same axes, plot graphs of vapour pressure of the mixture and of each component against temperature.
- Determine the boiling point of the mixture
- Calculate the percentage by mass of bromobenzene in the distillate.
- State how a sample of dry bromobenzene can be obtained from the distillate
- State one other application steam distillation.

3.a) Explain what is meant by the term.

i) Solubility product

ii) Common ion effect

b)(i) describe briefly how the solubility product of magnesium hydroxide in water can be determined.

ii) State how solubility product can be used to predict precipitation or dissolution.

c) The solubility product of calcium sulphate at 25°C is $2.4 \times 10^{-5}\text{mol}^2\text{dm}^{-6}$. Calculate the solubility of calcium sulphate in mol dm^{-3} in:

i) water

ii) 0.5M sulphuric acid

d) State one condition under which solubility product is invalid.

4. The table below shows the melting point of some oxides of elements in Group (IV) of the Periodic Table.

Compound	CO_2	SiO_2	GeO_2	SnO_2	PbO_2
Melting point ($^\circ\text{C}$)	-56.5	1700	116	1827	752

a) Explain the trend in the melting points of the oxides

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