

Answers to Sample Questions for the Chemistry Placement Exam

Note: These questions are intended to represent the general level of difficulty and areas of coverage of the Chemistry Placement Exam.

They are NOT intended to be a complete representation of the specific types of questions to be found on the exam.

1.	Aluminum reacts with element X to form a compound with the formula Al_2X_3 . Element X exists as diatomic molecules, X_2 , in the gaseous state at normal temperature and pressure. Element X must be
a)	Nitrogen
Answer b)	Oxygen
c)	Sulfur
d)	Chlorine
Explanation:	Aluminum forms a +3 ion. In order for the formula unit Al_2X_3 to be electrically neutral, the element X must form a -2 ion. Nonmetals in the oxygen column do this. Of these elements, only oxygen is a diatomic gas.
2.	What is the coefficient of O_2 when the following equation is balanced? $\text{C}_2\text{H}_6\text{O} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
a)	2
b)	3
c)	4
Answer d)	5
Explanation:	The balanced equation is $2 \text{C}_2\text{H}_6\text{O} + 5 \text{O}_2 \rightarrow 4 \text{CO}_2 + 4 \text{H}_2\text{O}$

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Craig Lane McFate



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