

### Types of Chemical Reactions Worksheet #1

#### **11-Chemistry**

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 1)  $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$  **Synthesis**
- 2)  $\text{CaI}_2 + \text{Cl}_2 \rightarrow \text{CaCl}_2 + \text{I}_2$  **Single Replacement**
- 3)  $\text{HNO}_3 + \text{AlCl}_3 \rightarrow \text{Al}(\text{OH})_3 + 3\text{HCl}$  **Double Replacement**
- 4)  $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$  **Synthesis**
- \*5)  $\text{Ca}(\text{OH})_2 + 2\text{HCl} \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$  **Double Replacement**
- 6)  $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$  **Decomposition**
- 7)  $3\text{CuSO}_4 + 2\text{Al} \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{Cu}$  **Single Replacement**
- 8)  $\text{Na}_2\text{S} + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{S}$  **Double Replacement**
- 9)  $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$  **Synthesis**

\*10)  $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$  **Single Replacement**

\*Remember: "H<sub>2</sub>O" can also be thought of as "HOH"

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- 11) Double replacement:  $\text{Na}_2\text{CrO}_4 + \text{PbCl}_2 \rightarrow 2\text{NaCl} + \text{PbCrO}_4$
- 12) Single replacement:  $\text{Cl}_2 + 2\text{NaBr} \rightarrow 2\text{NaCl} + \text{Br}_2$
- 13) Decomposition:  $\text{Mg}(\text{ClO}_3)_2 \rightarrow \text{MgCl}_2 + 3\text{O}_2$
- 14) Synthesis:  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- 15) Double replacement:  $3\text{Ca}(\text{OH})_2 + 2\text{FeCl}_3 \rightarrow 2\text{Fe}(\text{OH})_3 + 3\text{CaCl}_2$
- 16) Single replacement:  $\text{Fe} + \text{Cu}(\text{NO}_3)_2 \rightarrow \text{Fe}(\text{NO}_3)_2 + \text{Cu}$  [Assume Fe<sup>2+</sup>]
- 17) Decomposition:  $2\text{Hg}_2\text{O} \rightarrow 4\text{Hg} + \text{O}_2$
- 18) Synthesis:  $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$
- 19) Double replacement:  $\text{AgNO}_3 + \text{KI} \rightarrow \text{AgI} + \text{KNO}_3$
- 20) Single replacement:  $\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu}(\text{NO}_3)_2$  [Copper (II) is used here]

# Science Skills Chemical Reactions Answers

**Richard M. Gargiulo, Emily C.  
Bouck, SAGE Publications, Inc.**



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topics discussed include best practices project based education blended learning and the role of technology including e learning and science visualization Hands on recommendations on how to optimally implement innovative strategies of teaching chemistry at university and high school levels make this book an essential resource for anybody interested in either teaching or learning chemistry more effectively from experience chemistry professors to secondary school teachers from educators with no formal training in didactics to frustrated chemistry students

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