Forming Chemical Bonds Section 8.1

- I. Chemical Bonds
  - The force that holds two atoms together is called a chemical bond.
  - Chemical bonds may form by the attraction between a positive nucleus and negative electrons or the attraction between a positive ion and a negative ion.
  - c. Valence electrons are involved in the formation of chemical bonds between two atoms.
  - d. How do thousands of compounds form from the relatively few elements known to exist?
    - The answer lies in the electron structure of the atoms of the elements involved and the nature of the attractive forces between these atoms.
    - Recall that electron-dot structure is a type of diagram used to keep track of valence electrons and is especially useful when illustrating the formation of chemical bonds.
    - Recall that ionization energy refers to how easily an atom loses an electron. The term electron affinity indicates how much attraction an atom has for electrons.
    - iv. The difference in reactivity is directly related to the valence electrons.
- II. Formation of Positive Ions
  - Compare sodium and neon electron configuration.
  - b. Compare sodium ion and neutral sodium.
  - It is possible for atoms to achieve a pseudo-noble gas configuration.
    - Zinc electron configuration?
      - 1. What electrons would it lose?
  - d. A positively charged ion is called a cation.
- III. Formation of Negative Ions
  - Recall nonmetals have a great attraction for electrons and form a stable outer configuration by gaining electrons.
  - b. Chlorine electron configuration?
    - i. What happens?
    - ii. What noble gas does it become like?
  - c. A negatively charged ion is called an anion.

The Formation and Nature of Ionic Bonds Section 8.2

- I. Formation of an Ionic Bond
  - The electrostatic force that holds oppositely charged particles together in an ionic compound is referred to as an ionic bond.
- Compounds that contain ionic bonds are ionic compounds.
  - a. If ionic bonds occur between metals and nonmetal oxygen, oxides form.
  - b. Most other ionic compounds are called salts.
  - c. Binary ionic compounds contain a metallic cation and nonmetallic anion.
- III. What happens in the formation of a compound?
  - Explain the formation of the ionic compound composed of each pair of elements.
    - i. sodium and nitrogen

# **Section 8 1 Forming Chemical Bonds Answer Key**

**Angie Papple Johnston** 

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