

Stoichiometry

Quiz 11*9 Section 9

Section 9: Introduction to Stoichiometry

In the space provided, write the letter of the best response that best completes each problem or best answers each question.

1. Stoichiometry is the study of chemistry that deals with, drawing on concepts and facts that have been established in chemical reactions,
 - a. balancing
 - b. identifying
 - c. energy conversion
 - d. mass relationships
 - e. physical measurements
2. Some basic facts regarding the law of conservation of mass are
 - a. the law of conservation of mass.
 - b. that total mass is conserved.
 - c. that mass is conserved in all reactions.
 - d. temperature-related.
3. The number of significant figures in a product or quotient is usually the same as the number of significant figures in
 - a. the number of significant figures in the factor with the least number of significant figures.
 - b. the number of significant figures in the factor with the most significant figures.
 - c. the number of significant figures of the number with the most significant figures.
 - d. the number of significant figures in the number with the least number of significant figures.
4. The chemical equation $\text{P}_4\text{O}_{10} + 6\text{H}_2\text{O} \rightarrow 4\text{H}_3\text{PO}_4$ can be interpreted correctly as
 - a. 1 molecule of P_4O_{10} + 6 molecules of H_2O form 4 molecules of H_3PO_4 .
 - b. 1 mole of P_4O_{10} + 6 moles of H_2O form 4 moles of H_3PO_4 .
 - c. 1 g of P_4O_{10} + 6 g of H_2O form 4 g of H_3PO_4 .
 - d. both a and b.
5. How many moles of H_2SO_4 are produced from the following chemical equation?
 - a. 1
 - b. 2
 - c. 3
 - d. 4



Section Quiz Introduction To Stoichiometry

WJ Hussar



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