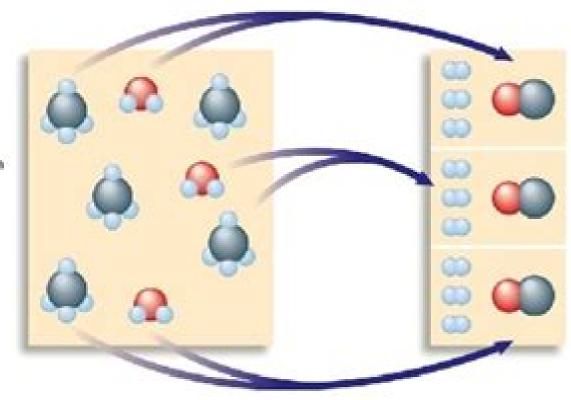
Limiting Reactants and Percent Yield

B. Calculations Involving a Limiting Reactant

A mixture of 5CH₄ and 3H₂O molecules undergoes the reaction CH₄(g) − H₂O(g) → 3H₂(q) + CO(q).

Note that the H₂O molecules are used up first, leaving two CH₄ molecules unreacted.



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