

1. In the list below, circle the properties of acids and box the properties of bases.

Sour taste	Feel slippery	Contains hydroxide ions
Turns litmus blue	Reacts with metals to form $H_2$	Turns litmus red
Bitter taste	Contains hydrogen ions	Corrosive

2. Complete the chart below:

Acid-Base Definitions		
Type	Acid Definition	Base Definition
Arrhenius	Donates an $H^+$ ion in solution	Donates an $OH^-$ ion in solution
Bronsted-Lowry	Donates an $H^+$ ion	Accepts an $H^+$ ion
Lewis	Accepts an electron pair	Donates an electron pair

3. What term(s) describe a substance that can act as an acid or a base? **Amphoteric or amphiprotic**
4. Neutralization reactions always produce a **SALT** and **WATER**
5. If the  $[H^+]$  is greater than the  $[OH^-]$ , the solution is acidic basic neutral
6. If the  $[OH^-]$  is greater than the  $[H^+]$ , the solution is acidic basic neutral
7. If the  $[H^+]$  is equal to the  $[OH^-]$ , the solution is acidic basic neutral
8. If the pH is 7, the solution is acidic basic neutral
9. If the pH is greater than 7, the solution is acidic basic neutral
10. If the pH is less than 7, the solution is acidic basic neutral
11. How many moles of lithium hydroxide are required to neutralize 19.2 grams of chloric acid?



$$(19.2 \text{ g } HClO_3) \times (1 \text{ mol } HClO_3 / 84 \text{ g } HClO_3) \times (1 \text{ mol } LiOH / 1 \text{ mol } HClO_3) = \underline{0.228 \text{ mol } LiOH}$$

12. How many grams of calcium hydroxide are required to neutralize 132.4 grams of nitric acid?



$$(132.4 \text{ g } HNO_3) \times (1 \text{ mol } HNO_3 / 63 \text{ g } HNO_3) \times (1 \text{ mol } Ca(OH)_2 / 2 \text{ mol } HNO_3) \times (74 \text{ g } Ca(OH)_2 / 1 \text{ mol } Ca(OH)_2) = \underline{77.76 \text{ g } Ca(OH)_2}$$

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**A Gutmann**



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## **Table of Contents Section 8 3 Properties Of Acids And Bases Worksheet Answers**

1. Understanding the eBook Section 8 3 Properties Of Acids And Bases Worksheet Answers
  - The Rise of Digital Reading Section 8 3 Properties Of Acids And Bases Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Section 8 3 Properties Of Acids And Bases Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Section 8 3 Properties Of Acids And Bases Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Section 8 3 Properties Of Acids And Bases Worksheet Answers
  - Personalized Recommendations
  - Section 8 3 Properties Of Acids And Bases Worksheet Answers User Reviews and Ratings

- Section 8 3 Properties Of Acids And Bases Worksheet Answers and Bestseller Lists
- 5. Accessing Section 8 3 Properties Of Acids And Bases Worksheet Answers Free and Paid eBooks
  - Section 8 3 Properties Of Acids And Bases Worksheet Answers Public Domain eBooks
  - Section 8 3 Properties Of Acids And Bases Worksheet Answers eBook Subscription Services
  - Section 8 3 Properties Of Acids And Bases Worksheet Answers Budget-Friendly Options
- 6. Navigating Section 8 3 Properties Of Acids And Bases Worksheet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Section 8 3 Properties Of Acids And Bases Worksheet Answers Compatibility with Devices
  - Section 8 3 Properties Of Acids And Bases Worksheet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Section 8 3 Properties Of Acids And Bases Worksheet Answers
  - Highlighting and Note-Taking Section 8 3 Properties Of Acids And Bases Worksheet Answers
  - Interactive Elements Section 8 3 Properties Of Acids And Bases Worksheet Answers
- 8. Staying Engaged with Section 8 3 Properties Of Acids And Bases Worksheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Section 8 3 Properties Of Acids And Bases Worksheet Answers
- 9. Balancing eBooks and Physical Books Section 8 3 Properties Of Acids And Bases Worksheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Section 8 3 Properties Of Acids And Bases Worksheet Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Section 8 3 Properties Of Acids And Bases Worksheet Answers
  - Setting Reading Goals Section 8 3 Properties Of Acids And Bases Worksheet Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 8 3 Properties Of Acids And Bases Worksheet Answers
  - Fact-Checking eBook Content of Section 8 3 Properties Of Acids And Bases Worksheet Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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