



12. Number of atoms in 560 g of Fe (atomic mass 56 g mol^{-1}) is :
 (1) Twice that in 70 g N (2) Half that in 20 g H
 (3) Both (1) and (2) (4) None of these
13. Which has maximum number of atoms :
 (1) 24 g of C (12) (2) 56 g of Fe (56) (3) 27 g of Al (27) (4) 108 g Ag (108)
14. The total number of protons, electrons and neutrons in 12 g of $^{12}_6\text{C}$ is :
 (1) 1.084×10^{25} (2) 6.022×10^{23} (3) 6.022×10^{22} (4) 18
15. 1.0 g of hydrogen contains 6×10^{23} atoms. The atomic weight of helium is 4. It follows that the number of atoms in 1 g of He is :
 (1) $\frac{1}{4} \times 6 \times 10^{23}$ (2) $4 \times 6 \times 10^{23}$ (3) 6×10^{23} (4) 12×10^{23}
16. A sample of aluminium has a mass of 54.0 g. What is the mass of the same number of magnesium atoms? (At. wt. Al = 27, Mg = 24)
 (1) 12 g (2) 24 g (3) 48 g (4) 96 g.
17. The number of atoms in 558.5 g of Fe (at wt. = 55.85) is :
 (1) Twice that in 60 g carbon (2) 6.022×10^{23}
 (3) Half in 8 g He (4) $558.5 \times 6.023 \times 10^{23}$
18. The total number of electrons present in 8.0 g of methane is.
 (1) 4.8×10^{24} (2) 3.01×10^{24} (3) 4.8×10^{23} (4) 3.01×10^{23} .
19. How many moles of electron weigh one kilogram :
 (1) 6.023×10^{23} (2) $\frac{1}{9.108} \times 10^{31}$ (3) $\frac{6.023}{9.108} \times 10^{24}$ (4) $\frac{1}{9.108 \times 6.023} \times 10^6$
20. Under the same conditions, two gases have the same number of molecules. They must
 (1) be noble gases (2) have equal volumes
 (3) have a volume of 22.4 dm^3 each (4) have an equal number of atoms
21. 5.6 litre of oxygen at STP contains :
 (1) 6.02×10^{23} atoms (2) 3.01×10^{23} atoms
 (3) 1.505×10^{23} atoms (4) 0.7525×10^{23} atoms
22. If V ml of the vapours of substance at NTP weight W g. Then molecular weight of substance is :
 (1) $(W/V) \times 22400$ (2) $\frac{V}{W} \times 22.4$ (3) $(W-V) \times 22400$ (4) $\frac{W \times 1}{V \times 22400}$
23. If 1 dm^3 of a gas weights 2.5 g at STP, its gram-molecular weight is :
 (1) 56 g (2) 11.2 g (3) 22.4 g (4) 224 g
24. 16 g of an ideal gas SO_x occupies 5.6 L, at STP. The value of x is
 (1) $x = 3$ (2) $x = 2$ (3) $x = 4$ (4) none of these

Section (C) : Density and Vapour density

1. Vapour density of a gas if its density is 0.178 g/L at NTP is :
 (1) 0.178 (2) 2 (3) 4 (4) 0.089
2. A gas is found to have the formula $(\text{CO})_x$. It's VD is 70 the value of x must be :
 (1) 7 (2) 4 (3) 5 (4) 6
3. Number of electrons in 1.8 mL of $\text{H}_2\text{O}(l)$ is about :
 (1) 6.02×10^{23} (2) 3.011×10^{23} (3) 0.6022×10^{21} (4) 60.22×10^{20}

Section (D) : Percentage composition and Molecular formula

1. The percentage value of nitrogen in urea is about
 (1) 46 (2) 85 (3) 18 (4) 28

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