

Name: _____

Date: _____

Student Exploration: Nuclear Decay

Vocabulary: alpha particle, atomic number, beta particle, daughter product, gamma ray, isotope, mass number, nuclear decay, positron, radioactive, subatomic particle

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

The chart below gives the locations, charges, and approximate masses of three **subatomic particles**. The approximate mass of each particle is given in universal mass units (u).

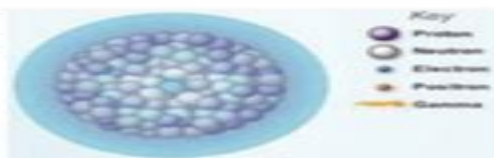
Particle	Location	Charge	Approximate mass
Proton	Nucleus	1^+	1 u
Neutron	Nucleus	0	1 u
Electron	Orbitals	1^-	0 u

- The **mass number** of an atom is equal to the sum of protons and neutrons in the nucleus. A helium atom has 2 protons and 2 neutrons. What is the mass number of this atom? _____
- The **atomic number** of an element is equal to the number of protons in each atom of the element. All helium atoms have 2 protons. What is the atomic number of helium? _____

Gizmo Warm-up

While most atoms are stable, some are **radioactive**, which means that they have a tendency to undergo spontaneous **nuclear decay**. The decay of radioactive atoms generally results in the emission of particles and/or energy.

Several types of nuclear decay can be explored with the Nuclear Decay Gizmo™. On the Gizmo, check that **Alpha decay** and **Uranium** are selected.



- Click **Play** (▶) and then click **Pause** (⏸) when the **alpha particle** is clearly visible. What is an alpha particle made of? _____
- Click **Play** and observe. Besides the alpha particle, what else is emitted from the nucleus during alpha decay? _____

Gamma rays are energetic electromagnetic waves; they are often emitted in nuclear decay.

Activity A: Alpha decay

Get the Gizmo ready:

- Click **Reset** (↺).
- Check that **Alpha decay** and **Uranium** are selected.



Question: How does alpha decay change the nucleus of a radioactive atom?

- Predict:** As you observed in the warm-up activity, an alpha particle consists of two protons and two neutrons. How will the emission of an alpha particle affect the following?
 - The atomic number of the atom: _____
 - The mass number of the atom: _____
- Calculate:** Turn on **Write equation**. What you see is an equation that shows the original uranium atom on the left. The boxes on the right represent the **daughter product**—the atom produced by radioactive decay—and the emitted alpha particle.
 - In the top left box, write the mass number of the daughter product and press "Enter" on your keyboard. What is this number? _____

Student Exploration Nuclear Decay Key

Iowa State University



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1. Understanding the eBook Student Exploration Nuclear Decay Key
 - The Rise of Digital Reading Student Exploration Nuclear Decay Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Student Exploration Nuclear Decay Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Student Exploration Nuclear Decay Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Student Exploration Nuclear Decay Key
 - Personalized Recommendations
 - Student Exploration Nuclear Decay Key User Reviews and Ratings
 - Student Exploration Nuclear Decay Key and Bestseller Lists
5. Accessing Student Exploration Nuclear Decay Key Free and Paid eBooks
 - Student Exploration Nuclear Decay Key Public Domain eBooks
 - Student Exploration Nuclear Decay Key eBook Subscription Services
 - Student Exploration Nuclear Decay Key Budget-Friendly Options
6. Navigating Student Exploration Nuclear Decay Key eBook Formats

- ePub, PDF, MOBI, and More
- Student Exploration Nuclear Decay Key Compatibility with Devices
- Student Exploration Nuclear Decay Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Student Exploration Nuclear Decay Key
 - Highlighting and Note-Taking Student Exploration Nuclear Decay Key
 - Interactive Elements Student Exploration Nuclear Decay Key
- 8. Staying Engaged with Student Exploration Nuclear Decay Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Student Exploration Nuclear Decay Key
- 9. Balancing eBooks and Physical Books Student Exploration Nuclear Decay Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Student Exploration Nuclear Decay Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Student Exploration Nuclear Decay Key
 - Setting Reading Goals Student Exploration Nuclear Decay Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Student Exploration Nuclear Decay Key
 - Fact-Checking eBook Content of Student Exploration Nuclear Decay Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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