

# Solubility and Temperature Gizmos

## Answer Key

You begin pouring sodium chloride into a glass of water. For a long time, the sodium chloride just dissolves in the water, but suddenly it begins to pile up at the bottom of the glass. Which statement is true?

- A. The sodium chloride has increased the temperature of the water.
- B. The sodium chloride has decreased the temperature of the water.
- C. The water is unsaturated.
- D. The water is saturated. - ANS D. The water is saturated.

Imagine you have two beakers. Both beakers are filled with the same amount of water. The water in both beakers is the same temperature as well. You add 50 g of Substance A to the first beaker, and 50 g of Substance B to the second beaker. After stirring both beakers, there is a small pile of Substance A at the bottom of the first beaker. None of Substance B is visible in the second beaker. Which of the following statements is true?

- A. Substance A is less soluble in water than Substance B.
- B. Substance A is more soluble in water than Substance B.
- C. Substance A is not soluble in water.
- D. Substance B is not soluble in water. - ANS Correct Answer: A. Substance A is less soluble in water than Substance B.

Which one of the following graphs best describes the solubility of potassium nitrate ( $\text{KNO}_3$ ) in water at different temperatures? - ANS Graph C

Comparing potassium nitrate ( $\text{KNO}_3$ ) and sodium chloride ( $\text{NaCl}$ ), which chemical's solubility (in water) is most affected by change in temperature?

- A. potassium nitrate
- B. sodium chloride
- C. same variation
- D. impossible to tell - ANS Correct Answer: A. potassium nitrate

Which of the charts below is most likely to represent a saturated solution? - ANS Correct Answer: C. Chart C

# Student Exploration Solubility And Temperature Answers Key

**Scott C. Dulebohn**



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