

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

**CHAPTER 4 STUDY GUIDE FOR CONTENT MASTERY****The Structure of the Atom****Section 4.1 Early Theories of Matter***In your textbook, read about the philosophers, John Dalton, and defining the atom.**For each statement below, write true or false.*

- false** \_\_\_\_\_ 1. Ancient philosophers regularly performed controlled experiments.
- true** \_\_\_\_\_ 2. Philosophers formulated explanations about the nature of matter based on their own experiences.
- true** \_\_\_\_\_ 3. Both Democritus and Dalton suggested that matter is made up of atoms.
- true** \_\_\_\_\_ 4. Dalton's atomic theory stated that atoms separate, combine, or rearrange in chemical reactions.
- false** \_\_\_\_\_ 5. Dalton's atomic theory stated that matter is mostly empty space.
- false** \_\_\_\_\_ 6. Dalton was correct in thinking that atoms could not be divided into smaller particles.
- true** \_\_\_\_\_ 7. Dalton's atomic theory stated that atoms of different elements combine in simple whole-number ratios to form compounds.
- true** \_\_\_\_\_ 8. Dalton thought that all atoms of a specific element have the same mass.
- false** \_\_\_\_\_ 9. Democritus proposed that atoms are held together by chemical bonds, but no one believed him.
- true** \_\_\_\_\_ 10. Dalton's atomic theory was based on careful measurements and extensive research.
- false** \_\_\_\_\_ 11. There are no instruments powerful enough to measure atoms so that they can be seen.
- true** \_\_\_\_\_ 12. The smallest particle of an element that retains the properties of that element is called an atom.

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**CHAPTER 4 STUDY GUIDE FOR CONTENT MASTERY****Section 4.2 Subatomic Particles and the Nuclear Atom***In your textbook, read about discovering the electron and the nuclear atom.**For each item in Column A, write the letter of the matching item in Column B.*

Column A	Column B
5. _____ 1. Proposed the nuclear atomic model	a. Thomson
6. _____ 2. Determined the mass-to-charge ratio of an electron	b. Millikan
7. _____ 3. Calculated the mass of an electron	c. Rutherford

*Draw and label a diagram of each atomic model.*

4. **plum pudding model**

Drawing should look like a ball of chocolate chip cookie dough. The chocolate chips should be labeled with negative charge or as electrons. The dough should be labeled as evenly distributed positive charges.

5. **no clear atomic model**

Drawing should look like a peach with a pit. The pit should be labeled nucleus and should include labeled protons and neutrons. The outer circle of the peach should be labeled electrons.

*In your textbook, read about the discovery of protons and neutrons.**Complete the following table of protons, electrons, and neutrons characteristics.*

Particle	Symbol	Location	Relative Charge	Relative Mass
1. Proton	$p^+$	In the nucleus	$1+$	1
2. Neutron	$n^0$	In the nucleus	0	1
3. Electron	$e^-$	In the space surrounding the nucleus	$1-$	$1/1840$

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# Study Guide Answers Chemistry Matter And Change

**ML Yell**



## **Study Guide Answers Chemistry Matter And Change:**

## Unveiling the Energy of Verbal Artistry: An Emotional Sojourn through **Study Guide Answers Chemistry Matter And Change**

In some sort of inundated with displays and the cacophony of instantaneous transmission, the profound energy and psychological resonance of verbal artistry frequently diminish into obscurity, eclipsed by the continuous barrage of sound and distractions. However, located within the musical pages of **Study Guide Answers Chemistry Matter And Change**, a interesting perform of literary elegance that pulses with organic thoughts, lies an wonderful trip waiting to be embarked upon. Written with a virtuoso wordsmith, that enchanting opus guides visitors on an emotional odyssey, softly exposing the latent potential and profound influence stuck within the intricate internet of language. Within the heart-wrenching expanse with this evocative evaluation, we will embark upon an introspective exploration of the book is key styles, dissect their captivating publishing fashion, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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