

## Physical Science Milestones Study Guide

**SPS1: Students will investigate our current understanding of the atom.**

**a. Examine the structure of the atom in terms of**

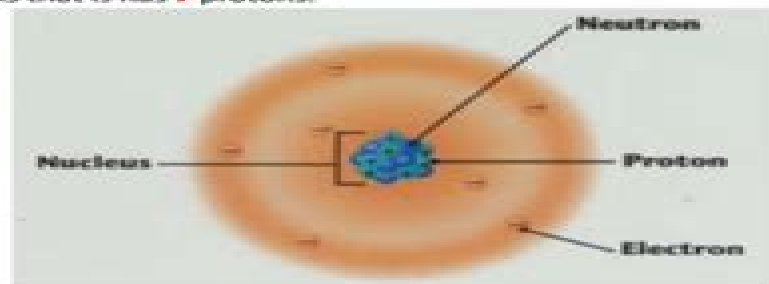
- proton, electron, and neutron locations.
  - **Proton** has a charge of **+1** and is located in the **nucleus** of the atom.
    - Protons are also the elements **Atomic** number and **define the identity of the element**, so protons are found by looking at the atomic number
  - Neutron has a charge of **0** and is located in the **nucleus** of the atom.
    - Neutrons are found by subtracting the **Atomic** Number from the **Atomic** Mass or the **BIG # - SMALL #**.
  - Electron has a charge of **-1** and is located in the **electron cloud** of the atom, outside of the nucleus.
    - Electrons are equal to the number of **Proton** sin a neutral atom.
  - The nucleus is made up of these 2 particles **Protons** and **Neutrons**.
  - The electron cloud is divided in to energy levels or energy shells, the first level will hold **2 electrons** electrons and the second level will hold **8** electrons.
  - The valence shell or valence level is the **Outer** energy shell of the atom. **The valence electrons determine the reactivity of an atom!!! Atoms will gain, lose, or share electrons in order to have a full valence of electrons which will make them more STABLE.**
  - In an uncharged atom, the number of **Protons** equals the number of **Electrons**.
  - Atoms, compounds and molecules have a **Neutral** charge.
  - The nucleus of an atom has a **Positive** charge.
- atomic mass and atomic number.
  - Atomic number is equal to the number of **Protons** in the atom and is also the atoms identity—it's the small number of the periodic table.
  - Atomic mass is equal to the average mass of all the isotopes of a particular element—it the big number on the periodic table. **Atomic Mass = Protons + Neutrons**
- atoms with different numbers of neutrons (isotopes). They also have a different atomic mass.
  - Ex. Carbon-12 and Carbon-14 **The -12 and -14 are the atomic MASSES.**
    - Carbon 12 has **6 Protons**    **6 Neutrons**    **6 Electrons**
    - Carbon 14 has **6 Protons**    **8 Neutrons**    **6 Electrons**
- explain the relationship of the proton number to the element's identity.
  - The **Atomic** number is equal to the number of **protons** in the atom which is how the element is identified.
  - Ex. Nitrogen has an atomic number of **7** which means that is has **7** protons.

molybdenum	← element name
42	← atomic number number of protons (Z)
Mo	← atomic symbol
95.94	← atomic mass A (this is an average mass)

- Number of Protons = Atomic Number
- Number of Electrons = Number of Protons = Atomic Number
- Number of Neutrons = Mass Number - Atomic Number

**b. Compare and contrast ionic and covalent bonds in terms of electron position.**

- Ionic bonds are the (sharing or **transferring**) of electrons and occur between a **metal** and a **non-metal**



# Study Guide Blank Physical Science

**JL Elias**



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