

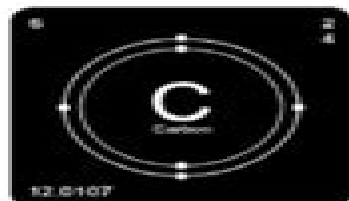
Study guide: Introduction to organic chemistry
Chemistry 12th

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Importance of Carbon



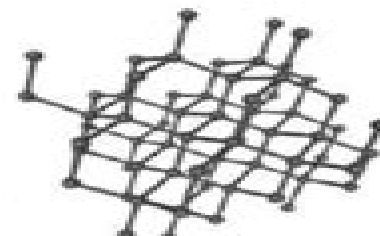
A compound found mainly in living things is known as an **organic compound**. **Organic compounds make up the cells and other structures of organisms and carry out life processes. Carbon is the main element in organic compounds**, so carbon is essential to life on Earth. Without carbon, life as we know it could not exist.

Why is carbon so basic to life? The reason is carbon's ability to form stable bonds with many elements, including itself, which is a rare property. This property allows carbon to form a huge variety of **very large and complex molecules**. In fact, there are nearly 10 million carbon-based compounds in living things! However, the millions of organic compounds can be grouped into just four major types: **carbohydrates, lipids, proteins, and nucleic acids (biochemical classification)**.

Natural Carbon sources

Pure Carbon occurs naturally with three main forms in nature:

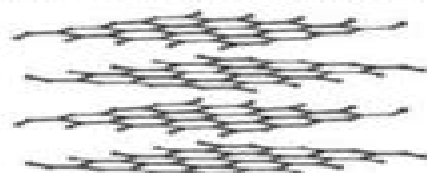
1. **Diamond:** Only carbon builds up this crystal. Each carbon atom binds to four other carbon atoms in a completely regular three-dimensional structure. Diamond is the hardest of all known minerals and has the highest fusion point at 3823 °C.



Application: jewelry, as abrasives in drill cores.

Formation: Formed under high pressure and temperature in the Earth's mantle, deep into the earth.

2. **Graphite:** Each carbon atom binds to three other carbon atoms in a hexagonal pattern. Each corner consists of a carbon atom. This pattern is spread as sheets and these sheets form layers. Weak bonds keep these sheets together. This means that graphite is one of the softer minerals on earth.

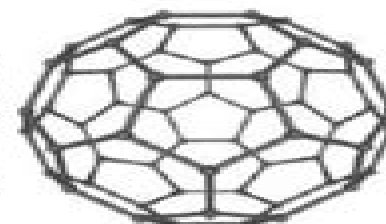


burn and thus be used as fuel.

Application: in pencils - the blacking is scraped off as layers that stick to the paper. Graphite conducts electricity poorly and is therefore used in electrical circuits as resistors. Graphite can

Formation: Graphite forms when plant parts and dead animals decompose in the absence of oxygen. This often happens while the material is packed together under sediment loads of eroded rock.

3. **Fullerene:** Each carbon atom binds to three other carbon atoms in patterns of hexagonal and pentagonal rings. In each corner is a carbon atom. These hexagonal and pentagonal structures build up forms that look like footballs and small tubes.



Application: Scientists are investigating the possibility to produce

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David R. Klein



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