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## CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY

### Solutions

#### Section 15.1 What are solutions?

In your notebook, read about the characteristics of solutions.

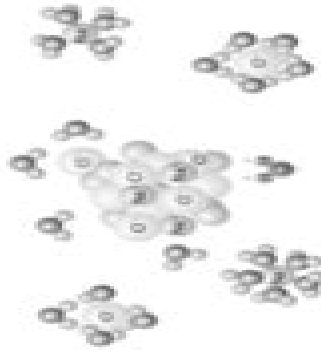
Use each of the terms below just once to complete the passage.

immiscible	liquid	soluble	solute	solution
miscible	insoluble	solvent		solvent

Air is an (1) solution of oxygen gas dissolved in nitrogen gas. The oxygen in air is the (2) solute and nitrogen is the (3) solvent. Because oxygen gas dissolves in a solvent, oxygen gas is an (4) soluble substance. A substance that does not dissolve in air is (5) insoluble. (6) Liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are (7) miscible. However, if one liquid is insoluble in another, the liquids are (8) immiscible.

Read about solvation in aqueous solutions in your textbook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



9. Hydration is solvation in which the solvent is water. What is solvation?  
**Solvation is the process of surrounding solute particles with solvent particles to form a solution.**

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## CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY

#### Section 15.1 continued

10. An sodium chloride dissolves in water. What happens to the sodium and chloride ions? **The sodium and chloride ions are separated and surrounded by the water molecules.**
11. Explain the orientation of the water molecules around the sodium ions and chloride ions. **Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.**
12. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? **How do you know?** **The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.**
13. List three ways that the rate of solvation may be increased. **stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent**

In your notebook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write *true* or *false*.

- true 14. The overall energy change that occurs when a solution forms is called the heat of solution.
- false 15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.
- false 16. Solvation continues as long as the solvation rate is less than the crystallization rate.
- true 17. In a saturated solution, solvation and crystallization are in equilibrium.
- true 18. Additional solute can be dissolved in an unsaturated solution.
- true 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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# Study Guide Content Mastery Water Resources

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