

# Chemical energy Study Guide

## Parent Signature (3 % extra credit on test)

I have gone over this study guide with my student for about 30 minutes.

### Chemical Vs. Physical

1. What is the definition of a chemical property?

*How a substance reacts with other substances. Can only be observed when a new substance is created.*

2. Name 3 chemical properties.

*Flammability, ability to rust, ability to fizz with acid, changes a strange color*

3. What is the definition of a physical property?

*Qualities of a substance that can be directly observed, creating a new substance is not necessary.*

4. Name 3 physical properties.

*Phase, ability to dissolve, mass, density, color, smell, texture*

5. Is the ability to dissolve a chemical or physical property?

*Physical—when something dissolves it does not change the substance, just breaks it into small pieces*

6. What is the difference between a chemical and a physical change?

*A chemical change creates a new substance, it is a chemical reaction. Physical changes don't create a new substance*

7. Give 3 examples of physical changes.

*Breaking something, melting or freezing, dissolving sugar*

8. Give 3 examples of chemical changes.

*Burning, fizzing, rusting, any chemical reaction*

### Chemical Reactions

1. Give 4 signs that can tell whether a chemical reaction has occurred (from your notes).

*Strange color change, strange change of phase, change in energy, change in smell/taste*

2. What is the definition of a reactant? Where do you find reactants in an equation?

*A chemical that makes a reaction happen, reactants are on the left side*

3. What is the definition of a product? Where do you find products in an equation?

*A chemical that is made from the reaction, products are on the right side*

# Study Guide Energy And Chemical Reactions Answers

**James E. Brady, Fred Senese**



## **Study Guide Energy And Chemical Reactions Answers:**

*Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual* Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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