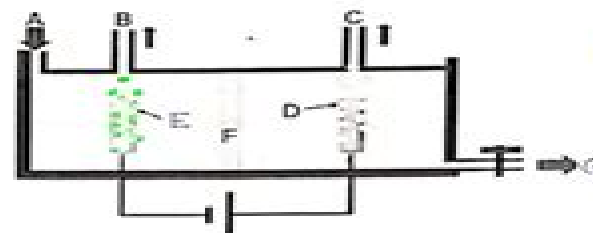


IGCSE – HONORS – ELECTROCHEMISTRY NOTES AND STUDY GUIDE:

Industrial electrolysis of Brine

Label each letter on the graph

- A. Burn in
- B. Hydrogen gas out
- C. Chloride gas out
- D. Anode
- E. Cathode
- F. Porous membrane
- G. Sodium hydroxide out



6. Electrolysis of Hydrochloric acid solution

Electrodes: Carbon (Graphite) (inert)

Electrolyte: Hydrochloric acid solution

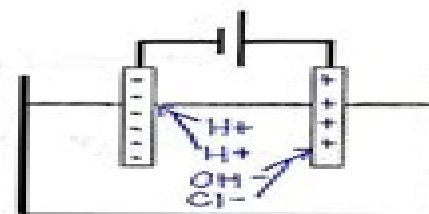
Ions Present: H^+ , OH^- , Cl^-

Half equations:

In the Cathode $2H^+ + 2e^- \rightarrow H_2$

In the Anode $2Cl^- \rightarrow 2e^- + Cl_2$

Overall Equation $2H^+ + 2Cl^- \rightarrow Cl_2 + H_2$



Observations:

Cathode: Bubbles of colorless gas (Hydrogen)

How can you prove it? A flame will pop

Anode: Bubbles of pale green gas (Chloride)

How can you prove it? Glowing

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