

## Chapter 6 Chemical Bonds

**Section 6.2 Covalent Bonding****(pages 166–169)**

This section discusses the formation of covalent bonds and the factors that determine whether a molecule is polar or nonpolar. It also discusses attractions between molecules.

**Reading Strategy** **page 166**

**Relating Text and Visuals** As you read the section, look closely at Figure 5. Complete the table for describing each type of model shown. For more information on this Reading Strategy, see the **Reading and Study Skills** in the Skills and Reference Handbook at the end of your textbook.

Molecular Models	
Model	Description
Electron dot	
Structural formula	
Space-filling	
Electron cloud	

**Covalent Bonds** **(pages 166–169)**

- Describe a covalent bond. \_\_\_\_\_  
\_\_\_\_\_
- Circle the letters of molecular models that show orbitals of atoms overlapping when a covalent bond forms.
  - electron dot
  - structural formula
  - space-filling
  - electron cloud
- Describe a molecule. \_\_\_\_\_  
\_\_\_\_\_
- Is the following sentence true or false? In a covalent bond, the atoms are held together by the attractions between the shared electrons and the protons in each nucleus. \_\_\_\_\_
- Circle the correct answer. Nitrogen has five valence electrons. How many pairs of electrons must two nitrogen atoms share in order for each atom to have eight valence electrons?
  - zero
  - one
  - two
  - three

## Section 6 2 Covalent Bonding Worksheet Answer Key

**N Colangelo**



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