

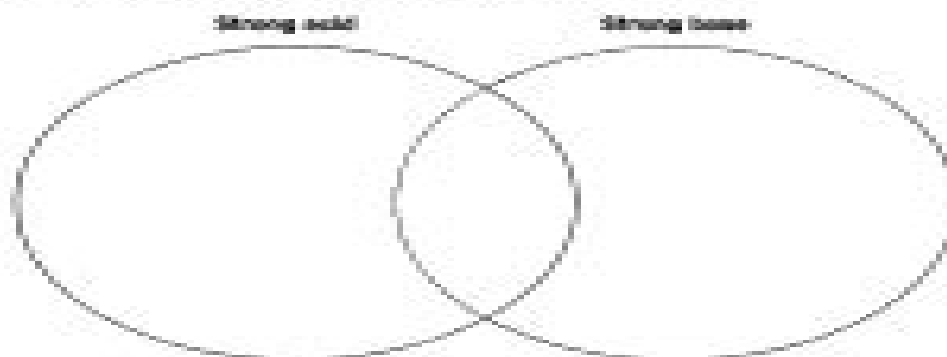
## Chapter 8 Solutions, Acids, and Bases

**Section 8.4 Strength of Acids and Bases****(pages 246–249)**

This section explains how to describe acids and bases in terms of both concentration and strength.

**Reading Strategy (page 246)**

**Comparing and Contrasting** As you read, complete the diagram by comparing and contrasting acids and bases. For more information on this reading strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

**The pH Scale (page 247)**

1. What is the name of the number scale chemists use to describe the concentration of hydronium ions in a solution?  
\_\_\_\_\_
2. The pH scale ranges from \_\_\_\_\_ to \_\_\_\_\_.
3. Circle the letter that indicates the pH of a neutral solution.  
a. 0  
b. 5  
c. 7  
d. 12
4. Water is neutral because it contains small but equal concentrations of \_\_\_\_\_ and \_\_\_\_\_.
5. Is the following sentence true or false? The higher the pH value of a solution, the greater the  $\text{H}_3\text{O}^+$  ion concentration is.  
\_\_\_\_\_
6. If you add acid to pure water, the concentration of  $\text{H}_3\text{O}^+$  \_\_\_\_\_ and the concentration of  $\text{OH}^-$  \_\_\_\_\_.

## Section 8 4 Strength Of Acids And Bases Answers

**DJ Losen**



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