Chemistry: Matter	and Chang	0	
Chapter Section:	8.1 & 8.2	Section Title:	Forming Chemical Bonds
Questions 1. Chemical bonds What are chemical bonds?		The force that holds two atoms together is called a chemical bond. Chemical bonds may form by the attraction between a positive nucleus and negative electrons or the attraction between a positive ion and a negative ion Electron dot structure is a type of diagram used to keep track of valence electrons and is especially useful when illustrating the formation of chemical bonds. A positively charged ion is called a cation. By losing an electron, the sodium atom acquired the stable outer electron configuration of neon, it is not neon. It is a sodium ion with a single positive charge. Transition metals have an outer energy level of ns going from left to right across a period, atoms of each element are filling an inner d sublevel. Although the formation of an octet is the most stable electron configuration. Other electron configurations provide some stability. Non metals located on the right side of the periodic table, have a great attraction for electrons and form stable outer electron configuration. Chlorine becomes an anion. Which is another name for a negative ion. To designate an anion, the ending -ide is added to the root name of the	
Properties of ionic compounds How are the positive and negative ions packed?		The strong attraction of positive ions and negative ions in an ionic compound results in a crystal lattice. melting point, boiling point, and hardness are physical properties that depend on how strongly the particles are attracted to each other. Energy changes also occur during the formation of ionic bonds from the ions formed during a chemical reaction. The energy required to separate one mole of the ions of an ionic compound is referred to as the lattice energy.	
 Energy and the io What is ionic bond? 	nic bond	Smaller ior The value The ionic b	rgy is directly related to the size of the ions bonded. Is generally have a more attraction for the valence electrons. If lattice energy is also affected by the charge of the ion. In ond formed from the attraction of ions with larger positive or sarged generally has a more negative lattice energy.

Section 8 Forming Chemical Bonds Answer Key

AW Rasmussen

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