9.1 THE ARITHMETIC OF EQUATIONS
SECTION REVIEW

Objectives

- Calculate the amount of reactants required or product formed in a nonchemical process.
- Interpret balanced chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

Key Terms

- stoichiometry

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

The coefficients of a balanced chemical equation indicate

the relative number of _____1_____ of reactants and products. All

stoichiometric calculations begin with a _____2_____. Only _____3_____

and _____4_____ are conserved in every reaction; moles, volumes,

and representative particles may not be.

In solving stoichiometric problems, conversion factors

relating moles of reactants to _____5_____ of products are used.

If you assume _____6_____, the equation also tells you about the

volume of gases.

1. _____

2. _____

3. _____

4. _____

5. _____

6. _____

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

_____ 7. The coefficients in a balanced chemical equation can be used to form mole ratios relating reactants to products.

_____ 8. The coefficients in a balanced chemical equation tell the relative volumes of reactants and products, expressed in any suitable unit of volume.

_____ 9. To calculate the mass of a molecule in grams, you can use the molar mass and Avogadro's number.

Section 9 1 The Arithmetic Of Equations

Joseph L. McCauley

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