


How Substances Dissolve

KEY IDEAS

As you read this section, keep these questions in mind:


- Why is water called the universal solvent?
- How do substances dissolve?

Why Do Substances Dissolve in Certain Solvents?

Water is often called a universal solvent because many substances dissolve in it. However, no one substance can dissolve every solute. A general rule in chemistry is that “like dissolves like.” This rule means that a solvent will dissolve substances that have similar molecular structures. 

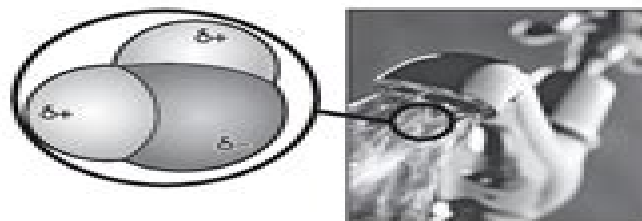
A solvent must be able to attract solute particles and pull them away from one another. Thus, solvent particles must attract solute particles more strongly than the solute particles attract one another.

POLAR MOLECULES

Recall that atoms bonded together do not all share electrons equally. For example, a water molecule is made up of two hydrogen atoms bonded to one oxygen atom. The oxygen atom attracts electrons more strongly than the hydrogen atoms do. Thus, the electrons spend more time near the oxygen atom. This gives each hydrogen atom a partial positive charge (δ^+) and the oxygen atom a partial negative charge (δ^-). 

Molecules that have partially charged positive and negative areas are **polar**. Because water molecules are polar, water can dissolve many other polar substances. Water can also dissolve ionic compounds such as table salt.

Water is a polar molecule because the oxygen atom strongly attracts electrons. Because the electrons spend more time near the oxygen atom, the oxygen atom has a partial negative charge.



READING TOOLBOX

Summarize Read each page of the section silently to yourself. With a partner, take turns summarizing each page.



READING CHECK

1. Identify Why is water called a universal solvent?

because many substances dissolve in it



READING CHECK

2. Explain Why do water molecules have partial charges?

The oxygen atom attracts electrons more strongly than the hydrogen atoms do. Thus, the electrons spend more time near the oxygen atom.

Section How Substances Dissolve Answer Key

Kevin T. Patton



Section How Substances Dissolve Answer Key:

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Scientifica David Sang, Peter Ellis, Derek McMonagle, 2004 Bring your science lessons to life with Scientifica Providing just the right proportion of reading versus doing these engaging resources are differentiated to support and challenge pupils of varying abilities [Cambridge Primary Science Stage 6 Teacher's Resource Book with CD-ROM](#) Fiona Baxter, Liz Dilley, 2014-05-22 Cambridge Primary Science is a flexible engaging course written specifically for the Cambridge Primary Science curriculum framework This Teacher's Resource for Stage 6 contains guidance on all components in the series Select activities and exercises to suit your teaching style and your learners abilities from the wide range of ideas presented Guidance includes suggestions for differentiation and assessment and supplementing your teaching with resources available online to help tailor your scheme of work according to your needs Answers to questions from the Learner's Book and Activity Book are also included The material is presented in editable format on CD ROM as well as in print to give you the opportunity to adapt it to your needs

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Chemistry: The Central Science Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Dalius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation

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Environmental Science Frank R. Spellman, Melissa L. Stoudt, 2013-02-14

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Table of Contents Section How Substances Dissolve Answer Key

1. Understanding the eBook Section How Substances Dissolve Answer Key
 - The Rise of Digital Reading Section How Substances Dissolve Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Section How Substances Dissolve Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section How Substances Dissolve Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section How Substances Dissolve Answer Key
 - Personalized Recommendations
 - Section How Substances Dissolve Answer Key User Reviews and Ratings

- Section How Substances Dissolve Answer Key and Bestseller Lists
- 5. Accessing Section How Substances Dissolve Answer Key Free and Paid eBooks
 - Section How Substances Dissolve Answer Key Public Domain eBooks
 - Section How Substances Dissolve Answer Key eBook Subscription Services
 - Section How Substances Dissolve Answer Key Budget-Friendly Options
- 6. Navigating Section How Substances Dissolve Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Section How Substances Dissolve Answer Key Compatibility with Devices
 - Section How Substances Dissolve Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section How Substances Dissolve Answer Key
 - Highlighting and Note-Taking Section How Substances Dissolve Answer Key
 - Interactive Elements Section How Substances Dissolve Answer Key
- 8. Staying Engaged with Section How Substances Dissolve Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section How Substances Dissolve Answer Key
- 9. Balancing eBooks and Physical Books Section How Substances Dissolve Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section How Substances Dissolve Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Section How Substances Dissolve Answer Key
 - Setting Reading Goals Section How Substances Dissolve Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section How Substances Dissolve Answer Key
 - Fact-Checking eBook Content of Section How Substances Dissolve Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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