Six Types of Chemical Reaction Worksheet Answers

Balance the following reactions and indicate which of the six types of chemical reaction are being represented:

-	AT THE PARTY OF TH	
ax)	2 NaBr + 1 Ca(OH) ₂ → 1 CaBr ₂ + 2 NaOH	Type of reaction: double displacement
(D)	2 NH ₂ + 1 H ₂ SO ₄ → 1 (NH ₄) ₂ SO ₄	Type of reaction: synthesis
c)	4 C ₂ H ₂ O + 27 O ₂ → 20 CO ₂ + 18 H ₂ O	Type of reaction: combustion
d)	3 Pb + 2 H ₃ PO ₄ -> 3 H ₂ + 1 Pb ₃ (PO ₄) ₂	Type of reaction: single displacement
(0)	1 Li ₃ N + 3 NH ₄ NO ₃ → 3 LINO ₃ + 1 (NH ₄) ₃ N	Type of reaction: double displacement
r)	3 HBr + 1 Al(OH)3 -> 3 H2O + 1 AlBr3	Type of reaction: acid-base (double displacement)
-90	Na ₃ PO ₄ + 3 KOH → 3 NaOH + K ₃ PO ₄	Reaction Typedouble displacement
h)	MgCl ₂ + Li ₂ CO ₃ → MgCO ₃ + 2 LiCl	Reaction Typedouble displacement
1)	C ₆ H ₁₂ + 9 O ₂ → 6 CO ₂ + 6 H ₂ O	Reaction Typecomplete combustion
(0)	Pb + FeSO₄ → PbSO₄ + Fe	Reaction Typesingle displacement
(k)	CaCO₂ → CaO + CO₂	Reaction Typedecomposition
10	P4 + 3 O2 → 2 P2O3	Reaction Typesynthesis
m) 2 RbNO ₃ + BeF ₂ → Be(NO ₃) ₂ + 2 RbF	Reaction Typedouble displacement
n()	2 AgNO ₅ + Cu → Cu(NO ₅) ₂ + 2 Ag	Reaction Typesingle displacement
0)	C ₃ H ₆ O + 4 O ₂ → 3 CO ₂ + 3 H ₂ O	Reaction Typecomplete combustion
p)	2 C ₅ H ₆ + Fe → Fe(C ₅ H ₆) ₂	Reaction Typesynthesis
93	SeCla + O₂ → SeO₂ + 3Cl₂	Reaction Typesingle displacement
(1)	$2 \text{ Mgl}_2 + \text{Mn}(SO_3)_2 \rightarrow 2 \text{ MgSO}_3 + \text{Mnl}_4$	Reaction Typedouble displacement
5)	O ₃ → O + O ₃	Reaction Typedecomposition
t)	2 NO₂ → 2 O₂ + N₂	Reaction Typedecomposition

Section Reactions And Equations Key

Darren J. Wilkinson

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