

Chapter 3 States of Matter

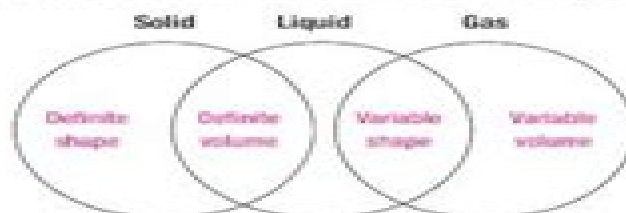
Section 3.1 Solids, Liquids, and Gases

(pages 68–73)

This section explains how materials are classified as solids, liquids, or gases. It also describes the behavior of these three states of matter.

Reading Strategy (page 68)

Comparing and Contrasting As you read about the states of matter, replace each letter in the diagram below with one of these phrases: *definite volume*, *definite shape*, *variable volume*, or *variable shape*. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

**Describing the States of Matter (pages 68–70)**

- What are three common states of matter?
a. Solids b. Liquids c. Gases
- Is the following sentence true or false? The fact that a copper wire can be bent shows that some solids do not have a definite shape.
false
- Circle the letter of each phrase that describes how particles at the atomic level are arranged within most solids.
a. randomly arranged ☒ b. packed close together
☒ c. arranged in a regular pattern d. spaced far apart
- Is the following sentence true or false? A liquid takes the shape of its container. true
- What is the state of matter in which a material has neither a definite shape nor a definite volume? gas
- Compare and contrast the arrangement of particles at the atomic level for a liquid and a solid. Particles in a solid are packed close together in an orderly arrangement. The arrangement of particles in a liquid is more random.
- What determines the shape and volume of a gas? A gas takes the shape and volume of its container.
- On the sun, where temperatures are extremely high, matter exists in a state known as plasma.

Section States Of Matter Answers

Montana. Supreme Court



Section States Of Matter Answers:

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web q1 what is the electron configuration of the element potassium answer the atomic number of potassium is 19 a neutral potassium element has 19 electrons its electron configuration is $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$ or simply put $4s^1$ q2 what is the total number of electrons that can be filled in s p and d subshells

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