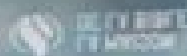
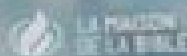


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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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$$8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$$

8.3 mole $\text{H}_2\text{S} \times 18 \text{ g } \text{H}_2\text{O} / \text{mole } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield
$$\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$$

problems 6 & 8 are similar to problem 5

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

if the reaction of 21.3 grams of O_2 produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have
$$\frac{21.3 \text{ g } \text{O}_2}{32 \text{ g/mole}} = 0.666 \text{ moles } \text{O}_2$$

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web a write a balanced equation for the reaction
$$4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$$

b determine the theoretical yield of Al_2O_3
$$4 \text{ mol Al} \times \frac{1 \text{ mol Al}_2\text{O}_3}{2 \text{ mol Al}} \times 101.96 \text{ g/mol Al}_2\text{O}_3 = 203.92 \text{ g Al}_2\text{O}_3$$

c determine the percent yield
$$\frac{160.0 \text{ g}}{203.92 \text{ g}} \times 100 = 78.5\%$$

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web 7Na + 2C₂O₄ is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be

weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h₂ are left over 13

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