



## Student Exploration: Stoichiometry

**Background:** A chemical equation is a shorthand representation of a chemical reaction. It shows the chemical formulas of the reactants and products, their relative amounts, and the conditions under which the reaction takes place.

Activity 1:  
Stoichiometry

Read the problem carefully.

1. Consider the equation below:



Problem

**Introduction:** You will use this equation to find out how much of each substance is needed to make a certain amount of product. This is called stoichiometry. It is a branch of chemistry that deals with the quantitative relationships between the amounts of reactants and products in a chemical reaction. The equation above shows that 1 mole of  $\text{Fe}_2\text{S}_3$  reacts with 6 moles of  $\text{HCl}$  to produce 2 moles of  $\text{FeCl}_3$  and 3 moles of  $\text{H}_2\text{S}$ . The equation is balanced, meaning that the number of atoms of each element is the same on both sides of the equation.



**Question:** How much of each substance is needed to make a certain amount of product? The equation above shows that 1 mole of  $\text{Fe}_2\text{S}_3$  reacts with 6 moles of  $\text{HCl}$  to produce 2 moles of  $\text{FeCl}_3$  and 3 moles of  $\text{H}_2\text{S}$ . The equation is balanced, meaning that the number of atoms of each element is the same on both sides of the equation.

**Question:** How much of each substance is needed?



- What is the ratio of the number of moles of  $\text{Fe}_2\text{S}_3$  to the number of moles of  $\text{HCl}$ ?
- What is the ratio of the number of moles of  $\text{Fe}_2\text{S}_3$  to the number of moles of  $\text{FeCl}_3$ ?
- What is the ratio of the number of moles of  $\text{Fe}_2\text{S}_3$  to the number of moles of  $\text{H}_2\text{S}$ ?
- What is the ratio of the number of moles of  $\text{Fe}_2\text{S}_3$  to the number of moles of  $\text{FeCl}_3$  and  $\text{H}_2\text{S}$ ?

**Question:** How much of each substance is needed to make a certain amount of product?

- Calculate the mass of each substance in the equation above. (The molar mass of  $\text{Fe}_2\text{S}_3$  is 204 g/mol, the molar mass of  $\text{HCl}$  is 36.5 g/mol, the molar mass of  $\text{FeCl}_3$  is 162.5 g/mol, and the molar mass of  $\text{H}_2\text{S}$  is 34 g/mol.)



# Student Exploration Stoichiometry Gizmo Answers Key

**CH Cherryholmes**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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