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Date: \_\_\_\_\_

## Student Exploration: Stoichiometry

**Vocabulary:** Avogadro's number, balanced equation, cancel, coefficient, conversion factor, dimensional analysis, molar mass, mole, molecular mass, stoichiometry

### Activity A: Moles

Get the Gizmo ready:

- Check that the equation is still:  
$$\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$$
- If not, click **New equation** until it reappears.



**Question:** How do scientists find the molecular mass and molar mass of a substance?

- Calculate:** The **molecular mass** of a molecule is the sum of the masses of each atom in the molecule. The unit of molecular mass is the universal mass unit (u).

Iron's atomic mass is 55.85 u, carbon's mass is 12.01 u, and oxygen's mass is 16.00 u.

- Calculate the molecular mass of carbon monoxide (CO) by adding the atomic mass of carbon and the atomic mass of oxygen: 28.01g
- Calculate the molecular mass of carbon dioxide (CO<sub>2</sub>): 44.01g
- Calculate the molecular mass of iron (III) oxide (Fe<sub>2</sub>O<sub>3</sub>): 159.7 g

- Infer:** A mole of a substance has a mass in grams that is equal to the molecular mass. For example, a carbon atom has a mass of 12.01 u. A mole of carbon has a mass of 12.01 g. Based on their molecular masses, list the **molar mass** of each substance:

Fe<sub>2</sub>O<sub>3</sub> 159.7 g    CO 28.01g    Fe 55.85g    CO<sub>2</sub> 44.01g

Check your answers on the Gizmo by inspecting the middle row of tiles on the right side of the Gizmo. These tiles show the units "1 mol" on top and "g" below.

- Practice:** Hydrogen has an atomic mass of 1.01 u. What is the molar mass of these substances?

H<sub>2</sub>O 18.02g    CH<sub>4</sub> 16.05g    H<sub>2</sub>CO<sub>3</sub> 62.03g    C<sub>6</sub>H<sub>12</sub>O<sub>6</sub> 180.18g

# Student Exploration Stoichiometry Worksheet Answers

**Rachel Sandford**



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