

General Student Growth Objectives



General Student Growth Objectives

There are two strategies to consider when setting a General SGO: simple and tiered.

- The **simple method** is based on determining how many of the total students are expected to meet a single target.
- The **tiered method** is based on expected growth within groups of students identified by their starting points, as discussed in SGO Step 2. (This is a part of the General SGO and should not be confused with a Specific SGO.)

Setting Simple Student Growth Objectives

To use the simple method of setting General SGOs, educators must predict what percentage of students would attain a particular level of performance on the final assessment. There are four levels of attainment of this objective. Shown in Figure 5 are the four levels of attainment possible for a student growth objective and what each level means.

Attainment of Student Growth Objective			
Exceptional 4	Full 3	Partial 2	Insufficient 1
Teacher has demonstrated an exceptional impact on learning by exceeding the objective.	Teacher has demonstrated a considerable impact on learning by meeting the objective.	Teacher has demonstrated some impact on learning but did not meet the objective.	Teacher has demonstrated an insufficient impact on learning by falling far short of the objective.

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Howard R. D. Gordon



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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable measure an appropriate volume of each liquid eg 25 cm³ place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with HCl as 57.9 kJ mol⁻¹

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic eg 1 HCl(aq) + NaOH(aq) → NaCl(aq) + H₂O(l) for which ΔH = 57 kJ mol⁻¹ eg 2 HNO₃(aq) + KOH(aq) → KNO₃(aq) + H₂O(l) for which ΔH = 57 kJ mol⁻¹

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