

Student Solution Manual

Student Solution Manual: Skoog et al., *Fundamentals of Analytical Chemistry*, 10e,
© 2022, 978-0-357-45055-0, Chapter 13: Complex Acid-Base Systems

Some of the answers below may differ in format but have the same value as your result. Please check with your instructor if a specific format is desired.

Chapter 13

- 13-1.** As its name implies, NaHA is an “acid salt” because it has a proton available to donate to a base. Briefly explain why a pH calculation for a solution of NaHA differs from that for a weak acid of the type HA.

Solution:

Not only is NaHA a proton donor, it is also the conjugate base of the parent acid H_2A .



Solutions of acid salts can be acidic or alkaline, depending on which of the above equilibria predominates. In order to calculate the pH of solutions of this type, it is necessary to take both equilibria into account.

- 13-4.** Why is it impossible to titrate all three protons of phosphoric acid in aqueous solution?

Answer:

The species HPO_4^{2-} is such a weak acid ($K_{a3} = 4.5 \times 10^{-13}$) that the change in pH in the vicinity of the third equivalence point is too small to be observable.

- 13-5.** Indicate whether an aqueous solution of the following compounds is acidic, neutral, or basic. Explain your answer.

Solutions:

(a) NaH_2PO_4



Solution will be acidic

(c) $NaNO_3$

Neither Na^+ nor NO_3^- reacts with H_2O .

Solution will be neutral

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Harvey Lodish



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