A system has reached equilibrium when: a) maximum entropy has been achieved b) minimum enthalpy has been achieved c) the rate of the forward reaction and reverse reaction is zero d) the concentrations of reactants and products have stopped changing Your answer is Explain why , conc vo Equilibrium is achieved when reactant and product concentrations (equal/constant/zero) 28. In a particular chemical reaction, $\Delta H = +100$ kJ. When equilibrium has been established; it is found that a significant amount of product has formed, even though there is still some reactants left. What has happened to entropy as this reaction was taking place? INVERSE C Explain bow you arrived at your answer FX. Wendothermic (AH 0005.) Given the following potential energy diagram for a reaction: 250 200 Products 150 PE (kJ) AH 6 pos 100

Progress of Reaction

Reactante

50

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Explain in terms of enthalpy and entropy, how you could end up with a fairly high patio of products to reactants. entre (AH pas) tend to min. enthalpy taxours reactants of the contract of th

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