

# Atomic Structure Review

Name \_\_\_\_\_

## 1. The 3 particles of the atom are:

	Location in Atom?	Size of LOCATION?	Charge?	Mass?
Protons	nucleus	.0001 %	+	1 amu
Neutrons	nucleus	.0001 %	neutral	1 amu
Electrons	cloud	99.999 %	-	.00054 amu

2. The number of protons in one atom of an element determines the atom's <sup>properties</sup> ~~charge~~ ~~#~~, and the number of electrons determines the charge of an element.

3. The atomic number tells you the number of protons in one atom of an element. It also tells you the number of electrons in a neutral atom of that element.

4. The atomic number gives the "identity" of an element as well as its location on the Periodic Table. No two different elements will have the same atomic number.

5. The avg atomic mass of an element is the average mass of an element's naturally occurring atom, or isotopes, taking into account the # of neutrons of each isotope.

6. The mass # of an element is the total number of protons and neutrons in the nucleus of the atom.

7. The mass number is used to calculate the number of neutrons in one atom of an element. In order to calculate the number of neutrons you must subtract the protons from the neutrons.

# Study Guide Answer The Atom Review

**David R. Klein, Laurie S. Starkey**



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