

Ex 1. For which of the following reactions would the  $\Delta H^\circ$  and the equilibrium constant  $K_{eq}$  be

1)  $\Delta H^\circ < 0$ ,  $K_{eq} > 1$   $\rightarrow$  **spontaneous**

2)  $\Delta H^\circ < 0$ ,  $K_{eq} < 1$   $\rightarrow$  **non-spontaneous**

3)  $\Delta H^\circ > 0$ ,  $K_{eq} > 1$   $\rightarrow$  **spontaneous**

4)  $\Delta H^\circ > 0$ ,  $K_{eq} < 1$   $\rightarrow$  **non-spontaneous**

Ex 2. The  $\Delta H^\circ$  for the reaction  $2\text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}_2(l) + \text{H}_2(g)$  is  $+136.3 \text{ kJ}$

The following reactions are spontaneous

$\text{H}_2\text{O}_2(l) \rightarrow \text{H}_2\text{O}(l) + \frac{1}{2}\text{O}_2(g)$   $\Delta H^\circ = -98.0 \text{ kJ}$

$\text{H}_2(g) + \frac{1}{2}\text{O}_2(g) \rightarrow \text{H}_2\text{O}(l)$   $\Delta H^\circ = -285.8 \text{ kJ}$

Calculate the  $\Delta H^\circ$  for the reaction of  $\text{H}_2\text{O}_2$  to  $\text{H}_2$  and  $\text{O}_2$

$\text{H}_2\text{O}_2(l) \rightarrow \text{H}_2(g) + \text{O}_2(g)$   $\Delta H^\circ = ?$

$\Delta H^\circ = 136.3 \text{ kJ}$   $\Delta H^\circ = -98.0 \text{ kJ}$   $\Delta H^\circ = -285.8 \text{ kJ}$

**Ex 3. Spontaneous** **Ex 4. non-spontaneous**

Ex 5. Calculate  $\Delta H^\circ$  for the following reaction at  $25^\circ\text{C}$

$\text{P}_4\text{O}_{10}(s) + 6\text{H}_2\text{O}(l) \rightarrow 4\text{H}_3\text{PO}_4(aq)$

$\Delta H^\circ_f \text{H}_3\text{PO}_4(aq) = -1279 \text{ kJ}$   $\Delta H^\circ_f \text{P}_4\text{O}_{10}(s) = -2942 \text{ kJ}$   $\Delta H^\circ_f \text{H}_2\text{O}(l) = -285.8 \text{ kJ}$

$\Delta H^\circ = 1279 \text{ kJ}$   $\Delta H^\circ = -2942 \text{ kJ}$   **$\Delta H^\circ = 2942 \text{ kJ}$**   $\Delta H^\circ = -285.8 \text{ kJ}$   $\Delta H^\circ = 1279 \text{ kJ}$

Ex 6. Calculate the heat of formation of  $\text{H}_2\text{O}(g)$  from the combustion data given below

$\text{H}_2(g) + \frac{1}{2}\text{O}_2(g) \rightarrow \text{H}_2\text{O}(l)$

**Heat of formation**

$\Delta H^\circ_f \text{H}_2\text{O}(l) = -285.8 \text{ kJ}$

$\Delta H^\circ_f \text{H}_2\text{O}(g) = -241.8 \text{ kJ}$

**$\Delta H^\circ_f \text{H}_2\text{O}(l) = -285.8 \text{ kJ}$**   $\Delta H^\circ_f \text{H}_2\text{O}(g) = -241.8 \text{ kJ}$   $\Delta H^\circ_f \text{H}_2\text{O}(l) = -285.8 \text{ kJ}$   $\Delta H^\circ_f \text{H}_2\text{O}(g) = -241.8 \text{ kJ}$

Ex 7. Calculate the standard heat of formation  $\Delta H^\circ_f$  of  $\text{H}_2\text{O}(g)$  from the following information:

$\text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}(g)$   $\Delta H^\circ = +44.0 \text{ kJ}$

$\Delta H^\circ_f \text{H}_2\text{O}(l) = -285.8 \text{ kJ}$

$\Delta H^\circ_f \text{H}_2\text{O}(g) = -241.8 \text{ kJ}$

$\Delta H^\circ_f \text{H}_2\text{O}(l) = -285.8 \text{ kJ}$

**$\Delta H^\circ_f \text{H}_2\text{O}(l) = -285.8 \text{ kJ}$**   $\Delta H^\circ_f \text{H}_2\text{O}(g) = -241.8 \text{ kJ}$   $\Delta H^\circ_f \text{H}_2\text{O}(l) = -285.8 \text{ kJ}$   $\Delta H^\circ_f \text{H}_2\text{O}(g) = -241.8 \text{ kJ}$

# Study Guide Answers Calculating Enthalpy

**Clyde Metz**



## Study Guide Answers Calculating Enthalpy:

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**Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

Study Guide to Accompany Calculus for the Management, Life, and Social Sciences Clyde Metz, 1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences

*Organic Chemistry Study Guide with Solutions Manual* Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer

*Study Guide and Solutions Manual* Neil E. Schore, 2002-08-02

**Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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Konstantinos Papadopoulos, 2012-07-06      Study Guide & Solutions Manual to Accompany Organic Chemistry, Third Edition  
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